

You must show all your calculation set-ups to receive full credit. You must report all answers to the correct # of significant figures (sig. figs.).

1. Who used a cathode ray tube in his/her discovery of the electron?

- ☒ a. Robert Millikan
☐ b. Ernest Rutherford
☐ c. John Dalton
☒ d. J.J. Thomson

82
100

2. How many significant figures are in each of the following measurements?

- a. 0.0017 g 2 sig figs
 b. 1390 mL 3 sig figs
 c. 4.030 km 4 sig figs
 d. 200110 kg 5 sig figs

3. Which unit represents a microgram?

- ☒ a. mg
☒ b. μ g
☐ c. ng
☐ d. cg

4. The most unreactive group in the Periodic Table is/are the

- ☒ a. alkaline-earth metals
☐ b. metalloids
☐ c. halogens
☒ d. noble gases

5. The empirical formula of a compound of carbon and hydrogen is CH_2 . The molar mass is 28.06 g/mol. What is the molecular formula?

- ☒ a. C_2H_4
☐ b. C_3H_6
☐ c. C_4H_8
☐ d. C_5H_{10}

C: 12.01
 H: 1.01

$$28.06 = n(14.03)$$

$$n = 2$$

$$\frac{\text{C } 12.01}{\text{H } 2(1.01)}$$

$$14.03$$

6. Label each of the following elements as metal, nonmetal, or metalloid.

- a. chlorine nonmetal
 b. beryllium metal
 c. silicon metalloid
 d. boron metalloid

7. Write both the name and the symbol of the element in group 13 (3A) and period 3.

Aluminum (Al)

8. Calculate the percentage composition by mass (mass percent) of just the hydrogen in C_3H_8 .

$$\text{C}_3\text{H}_8 \rightarrow \text{total: } 44.11$$

$$\text{C: } 3(12.01) = 36.03 / 44.11 = 82\%$$

$$\text{H: } 8(1.01) = 8.08 / 44.11 = 18\%$$

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