

## U8 Rev WS pg 2.jpg



## Chemistry I Chapter 11 Review

## Chemical Reactions

**Section 3:** Predict the products of chemical reactions and balance. Indicate precipitates and aqueous solutions as known.

1. \_\_\_\_\_  $\text{Ag(s)}$  + \_\_\_\_\_  $\text{CuSO}_4\text{(aq)}$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
2. \_\_\_\_\_  $\text{NaI(aq)}$  + \_\_\_\_\_  $\text{CaCl}_2\text{(aq)}$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
3. \_\_\_\_\_  $\text{O}_2\text{(g)}$  + \_\_\_\_\_  $\text{H}_2\text{(g)}$   $\rightarrow$  \_\_\_\_\_
4. \_\_\_\_\_  $\text{HNO}_3\text{(aq)}$  + \_\_\_\_\_  $\text{Mn(OH)}_2$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
5. \_\_\_\_\_  $\text{AgNO}_2\text{(aq)}$  + \_\_\_\_\_  $\text{BaSO}_4\text{(aq)}$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
6. \_\_\_\_\_  $\text{HCN}$  + \_\_\_\_\_  $\text{CuSO}_4\text{(aq)}$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
7. \_\_\_\_\_  $\text{LiBr(aq)}$  + \_\_\_\_\_  $\text{Co(SO}_3)_2$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
8. \_\_\_\_\_  $\text{HNO}_3$  + \_\_\_\_\_  $\text{Fe(OH)}_3$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
9. \_\_\_\_\_  $\text{C}_5\text{H}_{12}$  + \_\_\_\_\_  $\text{O}_2\text{(g)}$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_
10. \_\_\_\_\_  $\text{AgNO}_3\text{(aq)}$  + \_\_\_\_\_  $\text{Cu(s)}$   $\rightarrow$  \_\_\_\_\_ + \_\_\_\_\_

Identify the type of reaction for those equations above.

1. \_\_\_\_\_
2. \_\_\_\_\_
3. \_\_\_\_\_
4. \_\_\_\_\_
5. \_\_\_\_\_
6. \_\_\_\_\_
7. \_\_\_\_\_
8. \_\_\_\_\_
9. \_\_\_\_\_
10. \_\_\_\_\_

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**David Kirk**



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