

Study Card: Mass-to-Mass Stoichiometry

Success at solving stoichiometry problems requires ...

1. Understanding of **coefficients** (mole ratios)
Coefficients are used for mole to mole conversions
2. Understanding of **molar mass** (mass/1 mole)
Molar mass allows for a gram \leftrightarrow mole conversion
3. Ability to use the **Factor Label Method** to convert units
Conversion factors are the tool for converting units

Three Step Problem

A \rightarrow B

1. Grams A \rightarrow Moles A
Use Molar Mass of A
2. Moles A \rightarrow Moles B
Use Coefficients
3. Moles B \rightarrow Grams B
Use Molar Mass of B

Method:

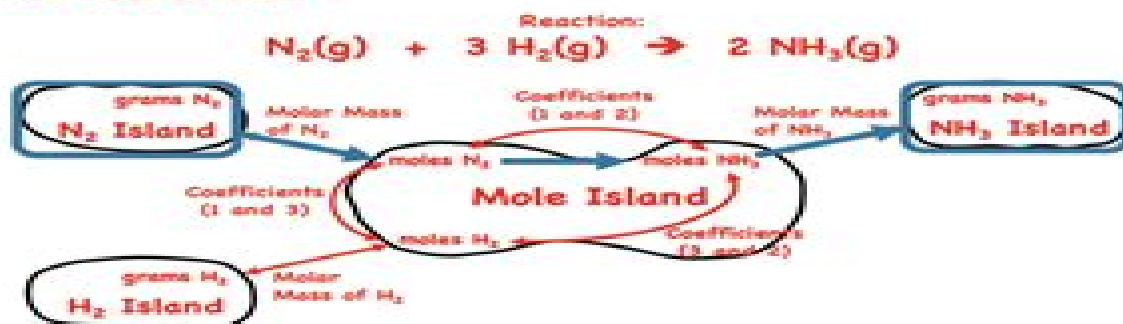
1. **Plan Your Solution**
Use Mole Island to organize your strategy.
2. **Set Up Conversion Factors** (Units; No #s)
Organize your (3) conversion factors to cancel units
3. **Insert Numbers**
Insert molar mass values and coefficients in front of units
4. **Calculate**
Use your calculator to calculate the answer.

Example 1



Determine the mass (in grams) of NH_3 produced from the reaction of 46.2 g of N_2 .

Plan Your Solution



Set Up Conversion Factors (Units, No #s)

$$46.2 \text{ g } \cancel{\text{N}_2} \cdot \frac{\text{mol } \cancel{\text{N}_2}}{\text{g } \cancel{\text{N}_2}} \cdot \frac{\text{mol } \text{NH}_3}{\text{mol } \cancel{\text{N}_2}} \cdot \frac{\text{g } \text{NH}_3}{\text{mol } \text{NH}_3} = \text{_____ g } \text{NH}_3$$

Insert numbers. Use calculator to solve.

$$46.2 \text{ g } \cancel{\text{N}_2} \cdot \frac{1 \text{ mol } \cancel{\text{N}_2}}{28.01 \text{ g } \cancel{\text{N}_2}} \cdot \frac{2 \text{ mol } \text{NH}_3}{1 \text{ mol } \cancel{\text{N}_2}} \cdot \frac{17.03 \text{ g } \text{NH}_3}{1 \text{ mol } \text{NH}_3} = \boxed{56.2 \text{ g } \text{NH}_3}$$

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Order is Important

Our observation is that the most frequent student frustrations occur when the early steps are skipped and students rush to their calculator without planning and setting up the solution. The calculator part should be easy. If you are finding it difficult, then you're probably skipping the first two or three steps of the **Method**.

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Robert R. Redfield



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