## 1 mole = $6.02 \times 10^{23}$ atoms 1 mole = atomic mass (g)



2. How many moles of barrum are in a cample containing 4.25 x 10" atoms of barrum?

#### Two Step Problems:

RACTICE PROBLEMS:

1. What is the mass of 2.25 x 10<sup>25</sup> atoms of lead?

1. How many motes are equal to 625g of crapper?

Convert 2:38 moles of carbon to alloms

4. How many atoms are in 4 Sg of pollassium?

2.25 x 1028 atoms of lead	1 mole	207.28	-	7744.19g = 7740g
	6.02 x 10 <sup>23</sup> atoms	1 mode		

2. How many atoms are in 10.0g of gold?

10 g gold	1 mole	$6.02 \times 10^{23}$ atoms	-	$3.06 \times 10^{22}$ atoms
	197.0e	1 mole		

# **Mass Mole Conversion Handout Answer Key**

Ying Liu

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