

F325 - June 2010, Q6e

17

- (e) Compound A is an orange ionic compound of chromium with the percentage composition by mass N, 11.11%; H, 3.17%; Cr, 41.27%; O, 44.45%. Compound A does not have water of crystallisation.

On gentle heating, compound A decomposes to form three products, B, C and water.

B is a green oxide of chromium with a molar mass of 152.0 g mol^{-1} .

C is a gas. At RTP, each cubic decimetre of C has a mass of 1.17 g.

In the steps below, show all your working.

- Calculate the empirical formula of compound A.
- Deduce the ions that make up the ionic compound A.
- Identify substances B and C.
- Write an equation for the decomposition of compound A by heat.

$$\begin{array}{r} \text{N} \\ 11.11\text{g} \\ \hline 14 \end{array}$$

$$\begin{array}{r} \text{H} \\ 3.17\text{g} \\ \hline 1 \end{array}$$

$$\begin{array}{r} \text{Cr} \\ 41.27\text{g} \\ \hline 52 \end{array}$$

$$\begin{array}{r} \text{O} \\ 44.45\text{g} \\ \hline 16 \end{array}$$

June 13 Chemistry F325

Robert R. Redfield



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