#### Question 2

- 2(a)(i) Comparison: X is more easily ionized than Mg.
  - Explanation: X is lower in group II than Mg. The outer electrons of X are in a shell further away from the nucleus than those of Mg. The outer electrons of X are therefore less

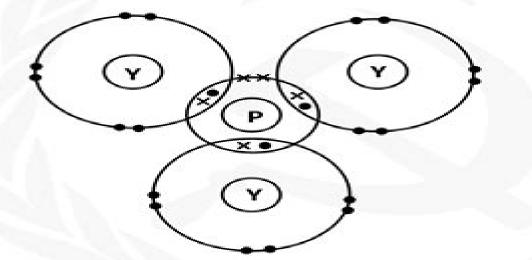
strongly held by the nucleus and more easily lost than the outer electrons of Mg.

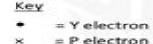
- 2(a)(ii) Comparison: Cl is a stronger oxidizing agent than Y.
  - Explanation: Oxidizing power of elements decreases down group VII. CI is higher in group VII than

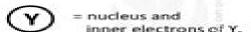
Y. The nucleus of CI is therefore less shielded than that of Y, making external

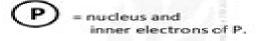
electrons more strongly attracted to the nucleus of CI than of Y.

- 2(b)(i) Both elements are solid at R.T.P.
- 2(b)(ii) At R.T.P. X conducts electricity but Y does not.
- 2(c)(i) Covalent bonding occurs between phosphorous and element Y.
- 2(c)(ii)









2(c)(iii) Chemical formula: PY3-

Two properties of compound: 1. It is a solid with a low melting point.

2. It is soluble in non-polar (organic) solvents.

## June 2013 F32ocr As Chemistry

**DP Hallahan** 

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#### Table of Contents June 2013 F32ocr As Chemistry

- 1. Understanding the eBook June 2013 F32ocr As Chemistry
  - The Rise of Digital Reading June 2013 F32ocr As Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying June 2013 F32ocr As Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an June 2013 F32ocr As Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from June 2013 F32ocr As Chemistry
  - Personalized Recommendations

- June 2013 F32ocr As Chemistry User Reviews and Ratings
- June 2013 F32ocr As Chemistry and Bestseller Lists
- 5. Accessing June 2013 F32ocr As Chemistry Free and Paid eBooks
  - June 2013 F32ocr As Chemistry Public Domain eBooks
  - June 2013 F32ocr As Chemistry eBook Subscription Services
  - June 2013 F32ocr As Chemistry Budget-Friendly Options
- 6. Navigating June 2013 F32ocr As Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - June 2013 F32ocr As Chemistry Compatibility with Devices
  - June 2013 F32ocr As Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of June 2013 F32ocr As Chemistry
  - Highlighting and Note-Taking June 2013 F32ocr As Chemistry
  - Interactive Elements June 2013 F32ocr As Chemistry
- 8. Staying Engaged with June 2013 F32ocr As Chemistry
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers June 2013 F32ocr As Chemistry
- 9. Balancing eBooks and Physical Books June 2013 F32ocr As Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection June 2013 F32ocr As Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine June 2013 F32ocr As Chemistry
  - Setting Reading Goals June 2013 F32ocr As Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of June 2013 F32ocr As Chemistry
  - Fact-Checking eBook Content of June 2013 F32ocr As Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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