



# Jeep Cherokee Power Seats Wiring

**Lauren Gardner**



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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h2o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{Fe}^{2+} + 5\text{H}_2\text{O} + \text{Cr}_2\text{O}_7^{2-} \rightarrow 4\text{Fe}^{3+} + 2\text{H}_2\text{O} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method  $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O} + 2\text{Cr}^{3+} + 9\text{H}_2\text{O}$   $2\text{N}_2\text{O}_4 + \text{O}_2 \rightarrow 2\text{Cr}^{3+} + 10\text{H}_2\text{O} + 4\text{Fe}^{3+} + 11\text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$  as  $\text{NO}_3^- + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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