

## Mixed Gas Laws Worksheet - Solutions

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11 \text{ moles of gas}$$

- 2) If 5.0 moles of  $\text{O}_2$  and 3.0 moles of  $\text{N}_2$  are placed in a 30.0 L tank at a temperature of  $25^\circ\text{C}$ , what will the pressure of the resulting mixture of gases be?  $25^\circ\text{C} = 298 \text{ K}$

$$\text{O}_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$\text{N}_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{Tot}} = P_{\text{O}_2} + P_{\text{N}_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is  $20.0^\circ\text{C}$ . By noon the temperature has risen to  $45.0^\circ\text{C}$ . What is the new volume of the balloon?

$$T_1 = 20.0^\circ\text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ\text{C} = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(315 \text{ K})} = 260 \text{ moles of gas}$$

- 5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in  $^\circ\text{C}$  of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})} = 296 \text{ K} = 23^\circ\text{C}$$

- 6)  $\text{CaCO}_3$  decomposes at  $1200^\circ\text{C}$  to form  $\text{CO}_2$  gas and  $\text{CaO}$ . If 25 L of  $\text{CO}_2$  are collected at  $1200^\circ\text{C}$ , what will the volume of this gas be after it cools to  $25^\circ\text{C}$ ?

$$T_1 = 1200^\circ\text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ\text{C} = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

- 7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at  $20.0^\circ\text{C}$  is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is  $-20.0^\circ\text{C}$ ?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ\text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ\text{C} = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

- 8) There are 135 L of gas in a container at a temperature of  $260^\circ\text{C}$ . If the gas was cooled until the volume decreased to 75 L, what would the temperature of the gas be?

$$T_1 = 260^\circ\text{C} = 533 \text{ K}, V_1 = 135 \text{ L}, T_2 = ?, V_2 = 75 \text{ L}$$

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(75 \text{ L})(533 \text{ K})}{(135 \text{ L})} = 296 \text{ K} = 23^\circ\text{C}$$

# Mixed Gas Laws Key

**L Darling-Hammond**



## **Mixed Gas Laws Key:**

*Mathematics and Science for Exercise and Sport* Craig Williams, David James, Cassie Wilson, 2008-06-30 This book is an introduction to the basic mathematical and scientific principles underpinning sport and exercise science. It is an invaluable course companion for students who have little prior experience of maths or science and an ideal revision aid for higher level undergraduate students. It is an ideal text for students of sport and exercise science, kinesiology and the human movement sciences.

*Phase Equilibria in Chemical Engineering* Stanley M. Walas, 2013-10-22 Phase Equilibria in Chemical Engineering is devoted to the thermodynamic basis and practical aspects of the calculation of equilibrium conditions of multiple phases that are pertinent to chemical engineering processes. Efforts have been made throughout the book to provide guidance to adequate theory and practice. The book begins with a long chapter on equations of state since it is intimately bound up with the development of thermodynamics. Following material on basic thermodynamics and nonidealities in terms of fugacities and activities, individual chapters are devoted to equilibria primarily between pairs of phases. A few topics that do not fit into these categories and for which the state of the art is not yet developed quantitatively have been relegated to a separate chapter. The chapter on chemical equilibria is pertinent since many processes involve simultaneous chemical and phase equilibria. Also included are chapters on the evaluation of enthalpy and entropy changes of nonideal substances and mixtures and on experimental methods. This book is intended as a reference and self study as well as a textbook either for full courses in phase equilibria or as a supplement to related courses in the chemical engineering curriculum. Practicing engineers concerned with separation technology and process design also may find the book useful.

*Science for Exercise and Sport* David James, Craig Williams, 2014-05-01 This handbook is written for undergraduate sport studies and sport and exercise students. It introduces students to the basic scientific principles that will underpin their learning and is aimed primarily at those who have little or no background in science. Craig Williams and David James apply key scientific concepts to real situations to better understand the principles at work. Clearly divided into three sections, the text covers the three physical states of gas, liquid and solid, explanations of forces, energy and electricity, including pressure, torque and joint velocity data analysis, ICT and report writing, important areas for the scientist. *Science for Exercise and Sport* provides the student with all the basic scientific background information they need and demonstrates how the theory can be used to map and monitor the human body in the sport and exercise discipline.

**Rosen's Emergency Medicine - Concepts and Clinical Practice E-Book** John Marx, Robert Hockberger, Ron Walls, 2013-08-01 Rely on Rosen's Emergency Medicine for the latest answers on every facet of emergency medicine practice. For decades this medical reference book has set the standard in emergency medicine, offering unparalleled comprehensiveness, clarity and authority to help you put the latest and best knowledge to work for your patients in the ER. Consult this title on your favorite e-reader, conduct rapid searches and adjust font sizes for optimal readability. Compatible with Kindle, nook and other popular devices. Practice confidently with ease.

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**The Undersea Journal** ,1997     Anesthesiology Keywords Review Raj K. Modak,2008 This review book is based on the keyword list distributed to residency program directors each year Each of the 316 entries begins with a keyword followed by a list of key points and a one or two page discussion written by a resident and reviewed by Yale University faculty Entries cover all areas of anesthesiology including pediatric anesthesia For easy reference the book is indexed in three ways alphabetically by rotation or organ system and according to the Boards organization A companion Website offers the fully searchable online text [www.keywordsreview.com](http://www.keywordsreview.com)

*Industrial Chemical Separation* Timothy C. Frank,Bruce S. Holden,2023-08-07 A fresh new treatment written by industry insiders this work gives readers a remarkably clear view into the world of chemical separation The authors review distillation extraction adsorption crystallization and the use of membranes providing historical perspective explaining key features and offering insights from personal experience The book is for engineers and chemists with current or future responsibility for chemical separation on a commercial scale in its design operation or improvement or for anyone wanting to learn more about chemical separation from an industrial point of view The result is a compelling survey of popular technologies and the profession one that brings the art and craft of chemical separation to life Ever wonder how popular separation technologies came about how a particular process functions or how mass transfer units differ from theoretical stages Or perhaps you want some pointers on how to begin solving a separation problem You will find clear explanations and valuable insights into these and other aspects of industrial practice in this refreshing new survey

Gas Industry ,1914     Public Utilities Reports Henry Clifford Spurr,Ellsworth Nichols,1922

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**Modern Universal Encyclopedia** Charles Higgins, Charles Annandale, R. Archer Johnson, H. D. Lovett, 1910 **Gas Age**, 1883 Includes summaries of proceedings and addresses of annual meetings of various gas associations L C set includes an index to these proceedings 1884 1902 issued as a supplement to Progressive age Feb 15 1910

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