

Name: _____ Date: _____ Pd: _____

MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.



STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: W5.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

$$\% \text{ Al} = \frac{\text{total mass Al}}{\text{molar mass Al}_2(\text{SO}_4)_3} \times 100 \% = \frac{53.96 \text{ g}}{342.17 \text{ g}} \times 100 \% = 15.77 \%$$

STEP 4: Complete the following problems using **a factor chart with units & chemicals**. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

1) How many **grams of calcium sulfate** are produced from reacting **2.59 moles of aluminum sulfate**?

$$2.59 \text{ mol Al}_2(\text{SO}_4)_3 \times \frac{3 \text{ mol CaSO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{136.15 \text{ g CaSO}_4}{1 \text{ mol CaSO}_4} = 1057.89 \text{ g CaSO}_4$$

moles A \longrightarrow moles B \longrightarrow grams B

Mixed Stoichiometry Answers

**James S. Mattson, Harry B.
Mark, Hubert C. MacDonald**



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Computer-assisted Instruction in Chemistry: General approach Harry B. Mark, Hubert C. MacDonald, 1974

Instructor's manual and solutions to problems to accompany General chemistry: principles and structures

James E. Brady, 1975-04 **Molecular Metals** William Hatfield, 2013-03-09 During the past few years there has been

intense research activity in the design synthesis and characterization of materials which are formed from molecular precursors and which have high or metal like electrical conductivities i e dcr dT *Merrill Chemistry* Robert C.

Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **The Practice of Chemistry** Donald J. Wink, Sharon

Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB **Scientific and**

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a sound physical understanding of the principles of analytical chemistry and their applications in the disciplines

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Rinehart and Winston Staff,2001 **Chemical Principles** Steven S. Zumdahl,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples **Chemical Process** Robin Smith,2005-06-10 This book deals with the design and integration of chemical processes emphasizing the conceptual issues that are fundamental to the creation of the process Chemical process design requires the selection of a series of processing steps and their integration to form a complete manufacturing system The text emphasizes both the design and selection of the steps as individual operations and their integration Also the process will normally operate as part of an integrated manufacturing site consisting of a number of processes serviced by a common utility system The design of utility systems has been dealt with in the text so that the interactions between processes and the utility system and interactions between different processes through the utility system can be exploited to maximize the performance of the site as a whole Chemical processing should form part of a sustainable industrial activity For chemical processing this means that processes should use raw materials as efficiently as is economic and practicable both to prevent the production of waste that can be environmentally harmful and to preserve the reserves of raw materials as much as possible Processes should use as little energy as economic and practicable both to prevent the build up of carbon dioxide in the atmosphere from burning fossil fuels and to preserve reserves of fossil fuels Water must also be consumed in sustainable quantities that do not cause deterioration in the quality of the water source and the long term quantity of the reserves Aqueous and atmospheric emissions must not be environmentally harmful and solid waste to landfill must be avoided Finally all aspects of chemical processing must feature good health and safety practice It is important for the designer to understand the limitations of the methods used in chemical process design The best way to understand the limitations is to understand the derivations of the equations used and the assumptions on which the equations are based Where practical the derivation of the design equations has been included in the text The book is intended to provide a practical guide to chemical process

design and integration for undergraduate and postgraduate students of chemical engineering practicing process designers and chemical engineers and applied chemists working in process development Examples have been included throughout the text Most of these examples do not require specialist software and can be performed on spreadsheet software Finally a number of exercises have been added at the end of each chapter to allow the reader to practice the calculation procedures

An Introductory Course of Quantitative Chemical Analysis Henry Paul Talbot, 1897

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