

Oxidation and Reduction

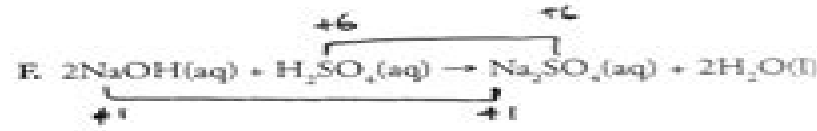
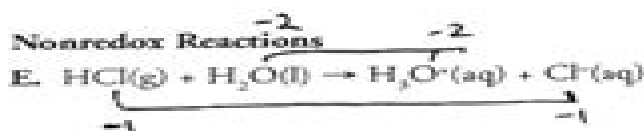
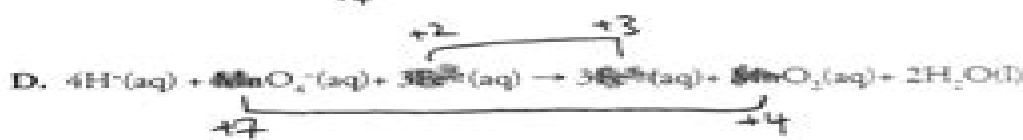
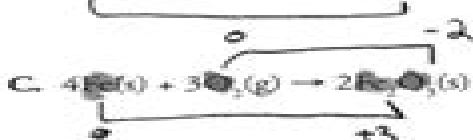
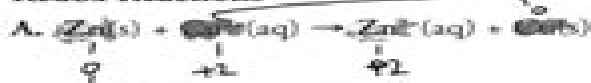
What happens when electrons are transferred in a chemical reaction?

Why?

Silver tarnishes when it comes in contact with sulfur compounds in the air. Copper gets coated in beautiful green patina as it ages. Metals rust or corrode in the presence of air and water. Minerals (ionic compounds) found in ore can be decomposed with the use of electricity to produce pure metals and nonmetals. All of these reactions are examples of oxidation and reduction, otherwise known as **redox** reactions. In this activity you will explore what is happening at the atomic level in redox reactions.

Model 1 – Redox Reactions

Redox Reactions



1. What two types of reactions are shown in Model 1?

Redox & Nonredox

2. Examine the redox and nonredox reactions in Model 1. Is/are there any feature(s) in the redox reactions that would allow you to identify them as redox reactions? If yes, use specific examples from Model 1 to support your answer.

Not necessarily

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A Loxley



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