

## Mixed Gas Laws Worksheet - Solutions

- 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?  

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11 \text{ moles of gas}$$
- 2) If 5.0 moles of  $\text{O}_2$  and 3.0 moles of  $\text{N}_2$  are placed in a 30.0 L tank at a temperature of  $25^\circ\text{C}$ , what will the pressure of the resulting mixture of gases be?  $25^\circ\text{C} = 298 \text{ K}$   

$$\text{O}_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$\text{N}_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{Tot}} = P_{\text{O}_2} + P_{\text{N}_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$
- 3) A balloon is filled with 35.0 L of helium in the morning when the temperature is  $20.0^\circ\text{C}$ . By noon the temperature has risen to  $45.0^\circ\text{C}$ . What is the new volume of the balloon?  

$$T_1 = 20.0^\circ\text{C} = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ\text{C} = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$
- 4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?  

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(315 \text{ K})} = 260 \text{ moles of gas}$$
- 5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in  $^\circ\text{C}$  of the balloon?  

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})} = 296 \text{ K} = 23^\circ\text{C}$$
- 6)  $\text{CaCO}_3$  decomposes at  $1200^\circ\text{C}$  to form  $\text{CO}_2$  gas and  $\text{CaO}$ . If 25 L of  $\text{CO}_2$  are collected at  $1200^\circ\text{C}$ , what will the volume of this gas be after it cools to  $25^\circ\text{C}$ ?  

$$T_1 = 1200^\circ\text{C} = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ\text{C} = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$
- 7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at  $20.0^\circ\text{C}$  is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is  $-20.0^\circ\text{C}$ ?  

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ\text{C} = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?,$$

$$T_2 = -20.0^\circ\text{C} = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$
- 8) There are 135 L of gas in a container at a temperature of  $260^\circ\text{C}$ . If the gas was cooled until the volume decreased to 75 L, what would the temperature of the gas be?  

$$T_1 = 260^\circ\text{C} = 533 \text{ K}, V_1 = 135 \text{ L}, T_2 = ?, V_2 = 75 \text{ L}$$

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(75 \text{ L})(533 \text{ K})}{(135 \text{ L})} = 296 \text{ K} = 23^\circ\text{C}$$

# Mixed Gas Law Worksheet Answers

**L Darling-Hammond**



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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Mixed Gas Law Worksheet Answers** a literary masterpiece penned by a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those who partake in its reading experience.

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