

## Chemistry Unit 8

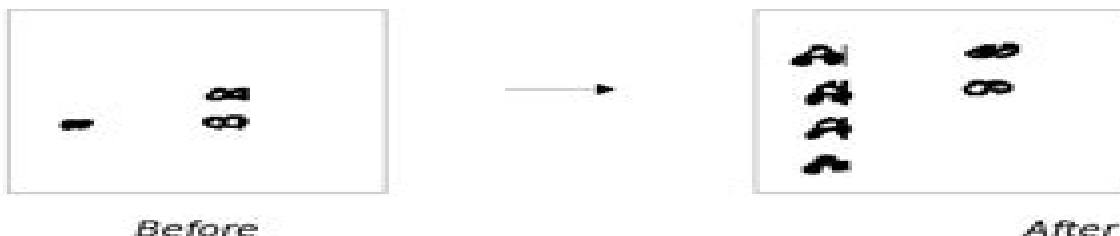
### Worksheet 3: Adjusting to Reality - Limiting Reactant

1. Write the balanced equation for the reaction between hydrogen and oxygen.

Balanced Equation:  $2\text{H}_2 + 1\text{O}_2 \rightarrow 2\text{H}_2\text{O}$

Suppose that 4 molecules of hydrogen gas and 4 molecules of oxygen gas react to form water.

Make a drawing that represents the reaction container before and after the reaction.



- 4 How many molecules of water can be produced?  
Oxygen, Leftover molecules Which reactant is in excess? Why?  
 2 How many molecules of excess reactant are there?

Construct a Before-Change-After Table for this reactant mixture:

**Bal. Equation:**  $2\text{H}_2 + 1\text{O}_2 \rightarrow 2\text{H}_2\text{O}$

<b>Before:</b>	<b>4</b>	<b>4</b>	<b>0</b>
<b>Change:</b>	<b>-4</b>	<b>-2</b>	<b>+4</b>
<b>After:</b>	<b>0</b>	<b>2</b>	<b>4</b>

According to the table you just made,

- 4 How many molecules of water can be produced?  
Oxygen, 2 molecules left over post-reaction Which reactant is in excess? Why?  
 2 How many molecules of excess reactant are there?

Based on your two methods of analysis above, what determines how much product can be made from a particular reactant mix?

The determining factor is which of the reactants runs out first. This determines how much product can be made, even if another reactant is in excess.

# Modeling Chemistry Unit 8 Quiz 3

**JE Gale**



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