Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO₃ (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| Equation: | 2 KClO ₃ → | 2 KCI + | 3 O2 |
|-----------|-----------------------|---------|------|
| Before: | 1.2 | 0 | 0 |
| Change: | -1.2 | +1.2 | +1.8 |
| After: | 0 | 1.2 | 1.8 |

1.2 moles KClO₃ needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| Equation: | Mig + | 2 HCl -> | MgCl ₂ + | H ₂ |
|-----------|--------|----------|---------------------|----------------|
| Before: | 0.100 | - 7 | 0 | 0 |
| Change: | -0.100 | -0.200 + | 0.160 + | 0.100 |
| Afters | 0 | | 0.100 | 0.100 |

0,200 moles HCI needed

0.73 g x 1 mole = .0040 mole 183 g

3. Ethane, C₂H₆ reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| Equation: | 2 C2H6 | + 7 O ₂ | → 4 CO₂ | + 6 H ₂ O |
|-----------|--------|--------------------|---------|----------------------|
| Before: | 2.20 | 36.56 | 0 | 0 |
| Change: | -2.20 | - 7.70 | + 4,40 | + 6.60 |
| After: | 0 | X 5 | 4.40 | 6.60 |

O₂ needed: 7.70 mole x 32.0 g = 246 g O₂ required 1 mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation: Ni(NO3)2 + 2 NaOH --> Ni(OH)2 + 2 NaNO3

 Equation:
 Ni(NO₃)₂ + 2 NaOH
 → Ni(OH)₂ + 2 NaNO₃

 Before:
 0.0040
 XS
 0
 0

 Change:
 -0.0040
 -0.0080
 +0.0040
 +0.0080

 After:
 0
 XS
 0.0040
 0.0080

NaNO, produced: 0.0080 mole x 85.0 g = 0.68 g 1 mole

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ML Yell

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