

## Unit 7: Balancing Chemical Reactions- Worksheet 2

Balance the following equations by inserting the proper coefficients.

1.  $\underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{SO}_3$  **Synthesis**
  2.  $\underline{2}\text{CH}_4 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO} + \underline{4}\text{H}_2\text{O}$  **Single R**
  3.  $\underline{2}\text{P} + \underline{3}\text{Cl}_2 \rightarrow \underline{2}\text{PCl}_3$  **Synthesis**
  4.  $\underline{2}\text{CO} + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CO}_2$  **Synthesis**
  5.  $\underline{2}\text{CH}_4 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{OH}$  **Synthesis**
  6.  $\underline{2}\text{Li} + \underline{1}\text{Br}_2 \rightarrow \underline{2}\text{LiBr}$  **Synthesis**
  7.  $\underline{2}\text{Al}_2\text{O}_3 \rightarrow \underline{4}\text{Al} + \underline{3}\text{O}_2$  **Decomp**
  8.  $\underline{2}\text{Na} + \underline{2}\text{H}_2\text{O} \rightarrow \underline{2}\text{NaOH} + \underline{1}\text{H}_2$  **Single R**
  9.  $\underline{6}\text{CO}_2 + \underline{6}\text{H}_2\text{O} \rightarrow \underline{1}\text{C}_6\text{H}_{12}\text{O}_6 + \underline{6}\text{O}_2$  **Rev. Comb**
  10.  $\underline{1}\text{H}_2\text{SO}_4 + \underline{2}\text{NaCl} \rightarrow \underline{2}\text{HCl} + \underline{1}\text{Na}_2\text{SO}_4$  **Double R**
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
11.  $\underline{3}\text{H}_2 + \underline{1}\text{SO}_2 \rightarrow \underline{1}\text{H}_2\text{S} + \underline{2}\text{H}_2\text{O}$  **Double R**
  12.  $\underline{2}\text{CaCO}_3 + \underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CaSO}_4 + \underline{2}\text{CO}_2$  **No**
  13.  $\underline{2}\text{AgNO}_3 + \underline{1}\text{CaCl}_2 \rightarrow \underline{2}\text{AgCl} + \underline{1}\text{Ca}(\text{NO}_3)_2$  **Double R**
  14.  $\underline{2}\text{HCl} + \underline{1}\text{Ba}(\text{OH})_2 \rightarrow \underline{1}\text{BaCl}_2 + \underline{2}\text{H}_2\text{O}$  **Double R**
  15.  $\underline{1}\text{H}_3\text{PO}_4 + \underline{3}\text{NaOH} \rightarrow \underline{1}\text{Na}_3\text{PO}_4 + \underline{3}\text{H}_2\text{O}$  **Double R**
  16.  $\underline{1}\text{Pb}(\text{NO}_3)_2 + \underline{2}\text{KI} \rightarrow \underline{1}\text{PbI}_2 + \underline{2}\text{KNO}_3$  **Double R**
  17.  $\underline{3}\text{CuO} + \underline{2}\text{NH}_3 \rightarrow \underline{1}\text{N}_2 + \underline{3}\text{Cu} + \underline{3}\text{H}_2\text{O}$  **Single R**
  18.  $\underline{1}\text{C}_2\text{H}_5\text{OH} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO}_2 + \underline{3}\text{H}_2\text{O}$  **Com**
  19.  $\underline{2}\text{C}_2\text{H}_6 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{COOH} + \underline{2}\text{H}_2\text{O}$  **No**
  20.  $\underline{3}\text{NO}_2 + \underline{1}\text{H}_2\text{O} \rightarrow \underline{2}\text{HNO}_3 + \underline{1}\text{NO}$  **No**

# Modeling Chemistry U7

**Camilla Rothe**



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In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Modeling Chemistry U7**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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