

Chemistry Unit 7 Review

701 – Inquiry Lab

1. Define...
 - a. Direct relationship: Two variables increase/decrease together
 - b. Indirect relationship: One variable increases, the other decreases
2. Identify whether the listed variables have a direct or indirect relationship to each other.
 - a. Moles and volume: Direct
 - b. Temperature and pressure: Direct
 - c. Pressure and volume: Indirect
 - d. Moles and pressure: Direct
3. Using the Ideal Gas Law equation, list 3 ways you could increase the volume of a balloon.
Increase the moles, increase the temperature, decrease the pressure (more so, the pressure drops as a result of greater volume)
4. Using the Ideal Gas Law equation, list 3 ways you could decrease the pressure within a syringe.
Decrease the moles, decrease the temperature, increase the volume
5. If the number of moles in a container is doubled, by how much should the pressure increase (assuming everything else is held constant)?
The pressure should also double

702 – Gas Laws Equation

1. What equation should you use when you are given just one set of variables?
 $PV = nRT$
2. You have a 4.5 L canister of O_2 that you store in a room that is $20\text{ }^{\circ}\text{C}$. The pressure of this tank is 13000 kPa. How many moles of O_2 are in this canister?
24 moles

Modeling Chemistry U7 Test Key

Rachel Sandford



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