

Chemistry Unit 8

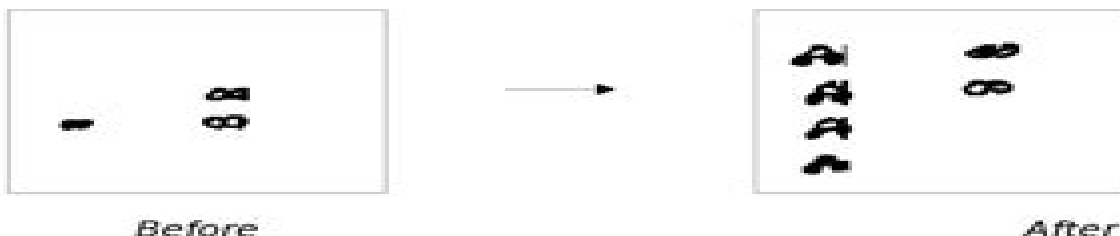
Worksheet 3: Adjusting to Reality - Limiting Reactant

1. Write the balanced equation for the reaction between hydrogen and oxygen.

Balanced Equation: $2\text{H}_2 + 1\text{O}_2 \rightarrow 2\text{H}_2\text{O}$

Suppose that 4 molecules of hydrogen gas and 4 molecules of oxygen gas react to form water.

Make a drawing that represents the reaction container before and after the reaction.



- 4 How many molecules of water can be produced?

Oxygen, Leftover molecules Which reactant is in excess? Why?

- 2 How many molecules of excess reactant are there?

Construct a Before-Change-After Table for this reactant mixture:

Bal. Equation: $2\text{H}_2 + 1\text{O}_2 \rightarrow 2\text{H}_2\text{O}$

| | | | |
|---------|----|----|----|
| Before: | 4 | 4 | 0 |
| Change: | -4 | -2 | +4 |
| After: | 0 | 2 | 4 |

According to the table you just made,

- 4 How many molecules of water can be produced?

Oxygen, 2 molecules left over post-reaction Which reactant is in excess? Why?

- 2 How many molecules of excess reactant are there?

Based on your two methods of analysis above, what determines how much product can be made from a particular reactant mix?

The determining factor is which of the reactants runs out first. This determines how much product can be made, even if another reactant is in excess.

Modeling Chemistry Unit 6 Test

J Rink



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