

**Unit 7 Chemical Reactions**  
**PRACTICE TEST/ REVIEW**

- \_\_\_\_\_ 1. A chemical reaction is a process in which
- a. products change into reactants
  - b. the law of conservation of mass applies
  - c. substances can change state
  - d. all of these
- \_\_\_\_\_ 2. During a chemical reaction,
- a. new elements are produced
  - b. atoms are rearranged
  - c. atoms are destroyed
  - d. elements are destroyed
- \_\_\_\_\_ 3. An equation is balanced by
- a. changing subscripts
  - b. adding coefficients
  - c. erasing elements as necessary
  - d. adding elements as necessary
- \_\_\_\_\_ 4. What are the reactants in the following chemical equation:
- $$\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$$
- a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 5. What are the products in the above equation?
- a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 6. The breaking down of a substance into two or simpler substances is \_\_\_\_\_
- a. Displacement
  - b. Synthesis
  - c. Decomposition
  - d. Double displacement
- \_\_\_\_\_ 7. Each substance to the right of the arrow in a chemical equation is a(n)
- a. Reactant
  - b. Catalyst
  - c. Product
  - d. Inhibitor
- \_\_\_\_\_ 8. According to the law of conservation of mass, if two atoms of hydrogen are used as a reactant, how many atoms of hydrogen must be part of the product?
- a. 4
  - b. 2
  - c. 1
  - d. 0
- \_\_\_\_\_ 9. What type of reaction is shown in the following chemical equation,  $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$ ?
- a. Displacement
  - b. Synthesis
  - c. Decomposition
  - d. Double displacement

# Modeling Chemistry Unit 7 Chemical Reactions Answer Key

**D Keegan**



## **Modeling Chemistry Unit 7 Chemical Reactions Answer Key:**

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**Chemical Reactions in Complex Mixtures** Ajit V. Sapre, Frederick J. Krambeck, 1991-07-25 In recent years there has been a convergence of trends in chemical reaction engineering and chemistry which have set the stage for significant advances in kinetic and thermodynamic modeling of processes New analytical chemistry methods new mathematical methods and new computational tools facilitate a more fundamental approach and a deeper understanding of chemical reactions in complex mixtures with very large numbers of compounds such as petroleum fractions This fortunate state of affairs has stimulated important new work both in academia and industrial research labs The purpose of the workshop that led to this book was to bring together researchers at the forefront of this field to review the state of the art stimulate communication and cooperation between industry and academia and develop a cohesive picture of research trends and future directions The chapters of the book have been organized into four main areas Continuous mixtures where the very large numbers of discrete compounds present are regarded as making up a continuum Structure activity relationships where the nature and rates of the reactions that a particular molecule undergoes are correlated with its chemical structure thus allowing the kinetics of very large numbers of compounds to be described by a few parameters Kinetic analysis where mathematical techniques are applied to analyze the behavior of kinetic networks and Thermodynamics emphasizing the practical and computational aspects of chemical equilibrium in complex mixtures

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