

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

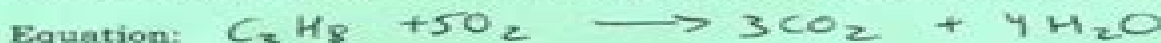


Before $\underline{8 \text{ moles}}$ $\xrightarrow{\text{excess}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-8 \text{ mol}}$ $\underline{-12 \text{ mol}}$ $\underline{+8 \text{ mol}}$ $\underline{+8 \text{ mol}}$

After — $\xrightarrow{\text{excess mol O}_2}$ $\underline{8 \text{ mol SO}_2}$ $\underline{8 \text{ mol H}_2\text{O}}$

2. Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?



Before $\underline{? \text{ mol}}$ $\xrightarrow{\text{excess}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-4 \text{ mol C}_3\text{H}_8}$ $\underline{-20 \text{ mol O}_2}$ $\xrightarrow{\quad}$ $\underline{12 \text{ mol CO}_2}$ $\underline{16 \text{ mol H}_2\text{O}}$

After $\underline{0 \text{ mol C}_3\text{H}_8}$ $\xrightarrow{\text{excess mol O}_2}$ $\underline{12 \text{ mol CO}_2}$ $\underline{16 \text{ mol H}_2\text{O}}$

$$12 \text{ mol CO}_2 \times \frac{5 \text{ mol O}_2}{3 \text{ mol CO}_2}$$

3. Ammonia, NH_3 , for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?



Before $\xrightarrow{\text{excess mol}}$ $\underline{0.15 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\xrightarrow{\text{excess H}_2}$ $\underline{-0.15 \text{ mol}}$ $\underline{0.3 \text{ mol}}$

After $\xrightarrow{\text{excess H}_2}$ $\underline{0 \text{ mol N}_2}$ $\underline{0.3 \text{ mol NH}_3}$

$$0.15 \text{ mol N}_2 \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} = 0.3 \text{ mol NH}_3$$

Modeling Chemistry Unit 8 1

Eric Strauss, Marylin Lisowski



Modeling Chemistry Unit 8 1:

Motor Vehicle-related Air Toxics Study ,1993 **Wetlands in Tooele Valley, Utah** Neil I. Burk,Charles Edwin Bishop,Mike Lowe,2005 This study indicates that wetlands in Tooele Valley are endangered The threats posed are from drought and increased development due to population growth which could dramatically affect the amount of water the wetlands receive sticker on back of disc cover The Summer Session ... Washington University (Saint Louis, Mo.),1927

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This book delves into Modeling Chemistry Unit 8 1. Modeling Chemistry Unit 8 1 is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Modeling Chemistry Unit 8 1, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Modeling Chemistry Unit 8 1
- Chapter 2: Essential Elements of Modeling Chemistry Unit 8 1
- Chapter 3: Modeling Chemistry Unit 8 1 in Everyday Life
- Chapter 4: Modeling Chemistry Unit 8 1 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Modeling Chemistry Unit 8 1. The first chapter will explore what Modeling Chemistry Unit 8 1 is, why Modeling Chemistry Unit 8 1 is vital, and how to effectively learn about Modeling Chemistry Unit 8 1.
3. In chapter 2, the author will delve into the foundational concepts of Modeling Chemistry Unit 8 1. This chapter will elucidate the essential principles that need to be understood to grasp Modeling Chemistry Unit 8 1 in its entirety.
4. In chapter 3, the author will examine the practical applications of Modeling Chemistry Unit 8 1 in daily life. The third chapter will showcase real-world examples of how Modeling Chemistry Unit 8 1 can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Modeling Chemistry Unit 8 1 in specific contexts. This chapter will explore how Modeling Chemistry Unit 8 1 is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Modeling Chemistry Unit 8 1. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Modeling Chemistry Unit 8 1.

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