

Lesson Plan in Campus Journalism

I. OBJECTIVES

At the end of the lesson, the students are expected to;

- A. Identify the reliable from unreliable source of news;
- B. Distinguish the different mediums of Journalism and;
- C. Write a short article about the news that you watched in television.

II. SUBJECT MATTER

Lesson: Identifying Features of Journalistic Writing

Reference: English 4

Materials: Power point Presentation

III. PROCEDURE

A. Pre-Activities

- Prayer
- Greetings
- Checking of Attendance

B. Motivational Activity

The teacher will display images of reliable and unreliable news sources. The students will determine whether or not the image displayed is reliable

1. Learning Task/Activity

C. Lesson Proper/Discussion

- **News** is the part of communication that keeps us informed of the changing events, issues, and characters in the world outside.
- The foremost value of news is as a utility to empower the informed.
- The purpose of journalism is thus to provide citizens with the information they need to make the best possible decisions about their lives, their communities, their communities, and their governments.
- **Journalism** is the act of gathering and presenting news and information to subjects for all mediums, print or non-print.
- The term journalism also refers to the news and information itself.
- There are different ways of presenting news and information including articles, reports, broadcasts, or even tweets.
- The journalist is always involved in the process of gathering selecting and presenting news and information being mindful of providing a truthful, noteworthy, and honest presentation. It should meet the standard of truth and honesty in reporting

Mediums of Journalism

1. Broadcast - reporters share the news through television or radio and through Video or audio online.
2. Print/Online - Journalist writes articles and blogs for publications like newspapers or magazines.
3. Photojournalism - Photojournalists tell stories through photos.
4. Multimedia - it is the combination of broadcast, print, and photojournalism.

Journalism Lesson Plans Worksheets

Lawrence A. Tomei



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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction e.g. particularly nitration of methyl benzoate to give methyl m-nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m-nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the

reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0 6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no2 which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1

h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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