

# Atomic Pudding Models of the Atom

Name \_\_\_\_\_

Date \_\_\_\_\_ Period \_\_\_\_\_

## Purpose

To compare various models of the atom that have appeared over the past two hundred years.

## Part 1: Using Evidence to Evaluate Models

In 1803, John Dalton proposed that atoms were simply solid spheres. In the decades that followed, chemists collected a lot of evidence suggesting that there was more to the atom. Look at the five models on the handout and read their descriptions. Then use the evidence in each box to evaluate the models.

**EVIDENCE 1:** It is possible to remove a negatively charged particle from an atom using electrical forces. (Thomson, 1897)

Does the evidence better support the solid sphere model or the plum pudding model? Explain your reasoning.

**EVIDENCE 2:** If a tiny particle is shot into the middle of an atom, it hits something dense in the center and bounces back in the direction from which it came. If a tiny particle is shot into the edges of the atom, it goes through. Most tiny particles shot at an atom will go through. (Rutherford, 1911)

Does the evidence better support the nuclear model or the plum pudding model? Explain your reasoning.

**EVIDENCE 3:** The farther from the center of an atom the negatively charged particles are, the easier they are to remove. (Bohr, 1913)

Does the evidence better support the solar system model or the plum pudding model? Explain your reasoning.

# Lesson 11 Atomic Pudding Answer Key

**N Colangelo**



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