

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{2 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?



$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mole Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$= \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?



$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{2 \text{ mol AgNO}_3} \times \frac{1 \text{ mole Cu}}{1 \text{ mole Cu}} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{2 \text{ mole KClO}_3} =$$

$$\underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed



$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{2 \text{ mol NaCl}} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mol Cl}_2} \times \frac{71 \text{ g Cl}_2}{1 \text{ mol Cl}_2}$$

$$= \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?



$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mol Cl}_2}{1 \text{ mole Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2}$$

$$= \underline{0.21 \text{ moles H}_2}$$

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Table of Contents Mhr Grade 12 Chemistry Practice Problems Answers

1. Understanding the eBook Mhr Grade 12 Chemistry Practice Problems Answers
 - The Rise of Digital Reading Mhr Grade 12 Chemistry Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mhr Grade 12 Chemistry Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mhr Grade 12 Chemistry Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mhr Grade 12 Chemistry Practice Problems Answers
 - Personalized Recommendations
 - Mhr Grade 12 Chemistry Practice Problems Answers User Reviews and Ratings
 - Mhr Grade 12 Chemistry Practice Problems Answers and Bestseller Lists

5. Accessing Mhr Grade 12 Chemistry Practice Problems Answers Free and Paid eBooks
 - Mhr Grade 12 Chemistry Practice Problems Answers Public Domain eBooks
 - Mhr Grade 12 Chemistry Practice Problems Answers eBook Subscription Services
 - Mhr Grade 12 Chemistry Practice Problems Answers Budget-Friendly Options
6. Navigating Mhr Grade 12 Chemistry Practice Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mhr Grade 12 Chemistry Practice Problems Answers Compatibility with Devices
 - Mhr Grade 12 Chemistry Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mhr Grade 12 Chemistry Practice Problems Answers
 - Highlighting and Note-Taking Mhr Grade 12 Chemistry Practice Problems Answers
 - Interactive Elements Mhr Grade 12 Chemistry Practice Problems Answers
8. Staying Engaged with Mhr Grade 12 Chemistry Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mhr Grade 12 Chemistry Practice Problems Answers
9. Balancing eBooks and Physical Books Mhr Grade 12 Chemistry Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mhr Grade 12 Chemistry Practice Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mhr Grade 12 Chemistry Practice Problems Answers
 - Setting Reading Goals Mhr Grade 12 Chemistry Practice Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mhr Grade 12 Chemistry Practice Problems Answers
 - Fact-Checking eBook Content of Mhr Grade 12 Chemistry Practice Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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