

Natural Science

Grade 8 Natural Science: Mid-Year Exam

Name _____ Date: 10 Minutes

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Instructions

- Read through the entire paper before answering the questions
- Answer all the questions on this question paper
- This paper consists of three sections: A, B and C

Section A: 10 marks

Section B: 10 marks

Section C: 10 marks

For the moderator

Section A	Section B	Section C
/ 10	/ 10	/ 10

Comments

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10

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True / False

True / False

True / False

True / False

True / False

True / False

10

Section A: 10 marks

10

Section B: 10 marks

Section C: 10 marks

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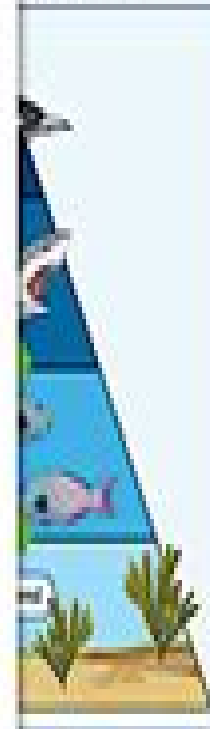
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Following question



Section A: 10 marks

Section B: 10 marks

Section C: 10 marks

Section A: 10 marks

Section B: 10 marks

Section C: 10 marks



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Section A: 10 marks

Section B: 10 marks

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Section A: 10 marks

Section B: 10 marks

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$
the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the