

Type 1 (Ionic) Bonding Worksheet

For each pair of elements below draw electron dot structures showing the valence electrons in each atom. Then draw arrows to show where the valence electrons will go during a chemical reaction. Write the name of each type I ionic compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Reactions- draw a picture showing each reaction & name the ionic compound	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{F} \Rightarrow \text{LiF}$ Name the ionic compound - <u>lithium fluoride</u>	Li	1	Li loses 1 e ⁻	Li^+
	F	7	F gains 1 e ⁻	F^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ Name the ionic compound - <u>calcium oxide</u>	Ca	2	Ca loses 2 e ⁻	Ca^{2+}
	O	6	O gains 2 e ⁻	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ Name the ionic compound - <u>Beryllium fluoride</u>	Be	2	loses 2 electrons	Be^{+2}
	F	7	gains 1 e ⁻	F^{-1}
	F	7	gains 1 e ⁻	F^{-1}
3) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ Name the ionic compound - <u>magnesium sulfide</u>	Mg	2	loses 2 e ⁻	Mg^{+2}
	S	6	gains 2 e ⁻	S^{-2}
4) $\text{K} + \text{F} \Rightarrow \text{KF}$ Name the ionic compound - <u>Potassium fluoride</u>	K	1	loses 1 e ⁻	K^{+1}
	F	7	gains 1 e ⁻	F^{-1}

Key Answer For Covalent Bond Notes Ch 8

John A. Olmsted, Gregory M. Williams



Key Answer For Covalent Bond Notes Ch 8:

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