2 A.M.U. Sci./Dip. Engg.		2007-2008		
(c) 350 cm, convergent		(c) Cl ₂ and CaSO ₄ (d) Cl ₂ and Ca(HCO ₃) ₂		
	The hand class is obtained by fusing:			
Three resistances A. B and C each of 50 and	-	A to the state of		
connected to a battery of 10 V as about to the		(b) a mixture of sand, lime, borax and arkan		
		(a) autosium carbonate and limescone		
W		(4) potentium early-posts, lead oxide and sand		
#v [⊥]	28.	Which of the following metals is most reactive?		
		(a) Aluminium (b) lead		
() 0		(c) Mercury (d) Silver Which of the following alloys contains		
	29.			
 The device used to generate electrical energy 		chromium? (a) Steel (b) Stainless steel		
	-	(c) Magnalium (d) Brass Which of the following is monomer of		
	30.	natural rubber?		
		(a) Chloroethene (b) Chloroprene		
(a) 0.25 X 10 ²³ (b) 0.50 X 10 ²³		(c) Isoprene (d) Buta-1, 3-diene		
		Alkaline KMnO ₄ oxidises propanone to:		
19. The value of charge / mass ratio of electron		(a) propanoic acid (b) ethanoic acid		
		(c) methanoic acid (d) oxalic acid		
(a) W.K. Roenteen (b) I.I. Thomson	22	Which of the following metals can displace Zn		
	34.	from ZnSO ₄ solution?		
		(a) Calcium (b) Copper		
example of combination reaction?		(c) Iron (d) Mercury		
(a) $Zn(s) + CuSO_{s}(sa) \rightarrow ZnSO_{s}(sa) + Cu(s)$	2.2	Heating of sodium ethanoate with soda lime		
(b) BaCl ₂ (aq) + Na ₂ SO ₂ (aq) → BaSO ₂ (s) + 2NaCl(aq)	33.	vields:		
$(c) CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$		(a) ethane (b) methane		
		(c) ethanol (d) methanol		
	2.4	The compound formed by the reaction of		
element "Y" to form compound with molecular		ethyne with bromine is:		
		(a) Br-CH=CH-Br (b) Br-CH ₂ -CHBr ₂		
The state of the s		(c) CH ₂ =CH-Br (d) Br ₂ CH-CHBr ₂		
	2.6	Cell were first discovered by Robert Hooke in		
	i altaba	the year:		
		(a) 1665 (b) 1674 (c) 1831 (d) 1835		
	24			
	30.	The plant cells have a rigid cell wall that lies: (a) Outside the plasma membrane		
(c) Li <na<k<rb (d)="" rb<k<na<li<="" td=""><td></td><td>A CAN TO THE PROPERTY OF THE P</td></na<k<rb>		A CAN TO THE PROPERTY OF THE P		
the state of the s		(b) Inside the plasma membrane		
		(c) In between the plasma membranes		
(a) Hydrogen (H) (b) Lithium (Li)	1900	(d) None of the above		
(c) Oxygen (O) (d) Chlorine (Cl)	37.	During mitosis nucleolus and nuclea		
Which of the following is an example of		membrane are lost in stage:		
strong electrolyte?		(a) Prophase (b) Metaphase		
(a) H ₂ CO ₁ (b) NH ₂ OH	25.00	(c) Anaphase (d) Telophase		
(c) NaCl (d) COOH - COOH	38.	An example of simple and permanent tissue is		
	1, 1414	(a) Xylem (b) Phloem		
		(C) Sciencia (d) All the above		
	39.	The shape of squamous epithelial tissue is:		
		(a) Cubical (b) Flattened		
		(c) Pillar like (d) None of the above		
	40.	Fresh water sponge belongs to phylum:		
	+	(a) Arthropoda (b) Annelida		
(a) CaCi2 and CaCO3 (b) Ci2 and Ca(On)2		(o) rumenda		
	Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C: SO	(c) 350 cm, convergent (d) 150 cm, divergent Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C: 28. (a) 2 amp (b) 1 amp 29. (c) 3 amp (d) 0.6 amp The device used to generate electrical energy is: (a) Electric motor (b) Generator (c) Galvanometer (d) Voltmeter The number of molecules in 11 g of CO ₂ are: (a) 0.25 X 10 ²⁹ (b) 0.50 X 10 ²⁹ (c) 1.00 X 10 ²⁹ (d) 1.51 X 10 ²⁹ 31. The value of charge / mass ratio of electron was determined by: (a) W K Roentgen (b) J.J.Thomson (c) Marie Curie (d) Niels Bohr Which of the following reactions is an example of combination reaction? (a) 2a(t) + Cu5O ₂ (a(t) + 2a5O ₂ (a(t) + Cu(t)) (b) 8c(3(t) + 3a5O ₂ (a(t) + BaSO ₂ (t) + 2hsC(t)a(t) (c) CaCO ₂ (t) → CO ₂ (t) (d) C(s) + Cu ₂ (t) + Co ₂ (t) (d) C(s) + Cu ₂ (t) + Cu ₂ (t) (e) CaCO ₂ (t) + Cu ₂ (t) (f) C(s) - Cu ₂ (t) + Cu ₂ (t) (f) C(s) - Cu ₂ (t) + Cu ₂ (t) (g) C(s) - Cu ₂ (t) + Cu ₂ (t) (h) XY (b) XY ₂ (c) XY ₁ (d) Xy ₂ (f) XY ₁ (d) Xy ₂ (g) XY ₁ (b) XY ₂ (g) XY ₂ (b) XY ₂ (g) XY ₃ (d) Xy ₃ The atomic radius (pm) of Li, Na, K and Rb varies in the order: (a) Na <k<rb<1 (b)="" k<na<li="">Rb<i (b)="" k<na<li="">Rb (c) Li<na<k<rb (d)="" (kj="" affinity="" electron="" mol<sup="" rb<k<na<li="" the="" value="">-1) of fluorine (F) is less than: (a) HyCO₂ (b) NH₄OH (c) NaCl (d) CoOH – COOH If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be: (a) two times (b) five times (c) ten times (d) twenty times Bleaching powder is manufactured by the</na<k<rb></i></k<rb<1>		

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Table of Contents Jipmer Entrance Exam Question Paper

- 1. Understanding the eBook Jipmer Entrance Exam Question Paper
 - The Rise of Digital Reading Jipmer Entrance Exam Question Paper
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Jipmer Entrance Exam Question Paper
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jipmer Entrance Exam Question Paper
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Jipmer Entrance Exam Question Paper
 - Personalized Recommendations
 - Jipmer Entrance Exam Question Paper User Reviews and Ratings
 - Jipmer Entrance Exam Question Paper and Bestseller Lists

- 5. Accessing Jipmer Entrance Exam Question Paper Free and Paid eBooks
 - o Jipmer Entrance Exam Question Paper Public Domain eBooks
 - Jipmer Entrance Exam Question Paper eBook Subscription Services
 - Jipmer Entrance Exam Question Paper Budget-Friendly Options
- 6. Navigating Jipmer Entrance Exam Question Paper eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Jipmer Entrance Exam Question Paper Compatibility with Devices
 - Jipmer Entrance Exam Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jipmer Entrance Exam Question Paper
 - Highlighting and Note-Taking Jipmer Entrance Exam Question Paper
 - Interactive Elements Jipmer Entrance Exam Question Paper
- 8. Staying Engaged with Jipmer Entrance Exam Question Paper
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jipmer Entrance Exam Question Paper
- 9. Balancing eBooks and Physical Books Jipmer Entrance Exam Question Paper
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Jipmer Entrance Exam Question Paper
- 10. Overcoming Reading Challenges
 - o Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Jipmer Entrance Exam Question Paper
 - Setting Reading Goals Jipmer Entrance Exam Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Jipmer Entrance Exam Question Paper
 - Fact-Checking eBook Content of Jipmer Entrance Exam Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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