OCR: Chemistry – Unit 2 – Chains, Energy and Resources

Note: For all Mechanisms and Reactions - There is an attached sheet at the end showing all the needed reactions in more detail.

THERE IS ALSO A LIST OF ALL EQUIATIONS AND REACTION NEEDED FOR THIS SYLLABUS ATTACHED IN A SEPARATE FILE.

(a) Interpret and use the terms:

 Empirical formula as the simplest whole number ratio of atoms of each element present in a compound.

 (ii) molecular formula as the actual number of atoms of each element in a molecule,

 (iii) general formula as the simplest algebraic formula of a member of a homologous series

| Alkane | C ₃ H ₂₅₊₂ | |
|-------------|-------------------------------------|---|
| Alkene | C _n H _{2n} | |
| Alcohol | C ₂ H _{2/22} OH | |
| Cycloalkane | C _n H _{2n} | |
| Cycloalkene | C _n H _{2n-2} | |
| Haloalkane | C.H.m.X | ======================================= |

(iv) structural formula as the minimal detail that shows the arrangement of atoms in a molecule

E.g. Butane:

СНЗСН2СН2СН3 or CH3(CH2)2CH3,

(v) displayed formula as the relative positioning of atoms and the bonds between them, ie for ethanol:

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M Walker

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