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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-H
Bond energy in kJ/mol	436	391

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an N-N bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4H_2) + 3(N_2)$$

$$4(436) + 3(165)$$

$$3150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond formed} = -90$$

$$3150 - x = -90$$

$$x = 3240$$

$$3240 \text{ kJ}$$
- Calculate the energy of an N-N bond.

$$6 \times (\text{N-H}) = 2346$$

$$\text{N-N} = \frac{3240 - 2346}{2}$$

$$447 \text{ kJ/mol}$$

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