

Acid-Base Test Review

Name _____

Directions: On a separate piece of paper, answer all questions in complete sentences or show all your work for full credit.

1. What is an acid?
 - a. What are the properties of acids
Differentiate between strong acids and weak acids, in terms of pH and $[H^+]$
2. What is a base?
 - a. What are the properties of bases
Differentiate between strong bases and weak bases, in terms of pH and $[OH^-]$
3. Explain what pH measures, and what it tells us about acids and bases.
4. What is the K_w ? What is the equation for K_w ?
5. Create a table showing the 3 acid-base definition.
6. Define and provide an example of the following:
 - a. Monoprotic –
 - b. Diprotic –
 - c. Triprotic –
7. What are conjugate acids and bases? Provide an example.
8. How is pH calculated?
9. Summarize naming rules for acids
10. Identify each property as applying to an acid, a base, or both.
 - a. bitter taste –
 - b. electrolyte –
 - c. indicator color change –
 - d. sour taste –
11. Write the formula for each acid or base
 - a. barium hydroxide –
 - b. hydrobromic acid –
 - c. rubidium hydroxide –
12. Name or give the molecular formula for each of the following acid or base.
 - a. Acetic acid-
 - b. Phosphorous acid-
 - c. $HClO_2$ -
 - d. H_3PO_4 -
 - e. sulfuric acid-
 - f. $HClO$ -
 - g. HF –
 - h. $HClO_3$ –
 - i. H_2CO_3 –
 - j. $Al(OH)_3$ –
13. Define acidic solution –
14. Define basic solution –
15. Define pH –
16. Explain how you would find the pH of a solution if you had the OH^- concentration.

Key Acid Base Study Guide

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