

Chemistry Unit 6 Worksheet 1: Intermolecular Forces

1. Predict the molecular shape of each of the following:

- a. H_2S _____ b. CCl_4 _____
c. SO_2 _____ d. PCl_5 _____

2. List all types of IMFs that would occur in each of the following (you should have a good enough understanding of electronegativities to answer all of these, except maybe SO_2 , without look up the electronegativity numbers).

- a. CH_3CF_3 _____
b. CCl_4 _____
c. SO_2 _____
d. BrF _____
e. $(\text{CH}_3)_3\text{N}$ _____
f. PCl_5 _____

3. H_2S , O_2 and CH_3OH all have comparable molecular masses. List the dominant type of IMF for the pure substances, then rank the strength of each compound based on IMFs within the samples.

4. Circle all of the species below that can form a hydrogen bond in its pure form. Explain why the other species couldn't hydrogen bond.



5. Rank the following compounds from weakest intermolecular forces to strongest. Justify your answers.



6. Rank the following from weakest intermolecular forces to strongest. Justify your answers.



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Steve DeRoy



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1. Understanding the eBook Living By Chemistry Unit Worksheet Answers 2006
 - The Rise of Digital Reading Living By Chemistry Unit Worksheet Answers 2006
 - Advantages of eBooks Over Traditional Books
2. Identifying Living By Chemistry Unit Worksheet Answers 2006
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Living By Chemistry Unit Worksheet Answers 2006
 - User-Friendly Interface
4. Exploring eBook Recommendations from Living By Chemistry Unit Worksheet Answers 2006
 - Personalized Recommendations
 - Living By Chemistry Unit Worksheet Answers 2006 User Reviews and Ratings
 - Living By Chemistry Unit Worksheet Answers 2006 and Bestseller Lists
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 - Living By Chemistry Unit Worksheet Answers 2006 Budget-Friendly Options
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 - Living By Chemistry Unit Worksheet Answers 2006 Enhanced eBook Features
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- Adjustable Fonts and Text Sizes of Living By Chemistry Unit Worksheet Answers 2006
 - Highlighting and Note-Taking Living By Chemistry Unit Worksheet Answers 2006
 - Interactive Elements Living By Chemistry Unit Worksheet Answers 2006
8. Staying Engaged with Living By Chemistry Unit Worksheet Answers 2006
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Living By Chemistry Unit Worksheet Answers 2006
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- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Living By Chemistry Unit Worksheet Answers 2006
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Living By Chemistry Unit Worksheet Answers 2006
- Setting Reading Goals Living By Chemistry Unit Worksheet Answers 2006
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Living By Chemistry Unit Worksheet Answers 2006
- Fact-Checking eBook Content of Living By Chemistry Unit Worksheet Answers 2006
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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