

## Review #6: Kinetic Molecular Theory of Matter and the Gas Laws

### 1. Definitions:

- a) **The Kinetic Molecular Theory (KMT)** is a theory that states that all matter is made up of tiny particles (either atoms or molecules) and that these particles are in constant motion.
- b) **Kinetic Energy** is the energy that objects have due to their motion. As the amount of motion of an object increases, kinetic energy also increases.
- c) **Potential Energy** is the energy that objects have due to their position. As the amount of space between two objects increases, their potential energy (relative to each other) also increases.
- d) **Temperature** is a measure of the average kinetic energy of the particles in a substance. The Kelvin scale is a measure of the absolute amount of motion the particles have, such that zero Kelvins (absolute zero) means zero motion of the particles.
- e) **STP** refers to Standard Temperature and Pressure, and this has been defined as exactly  $0^{\circ}\text{C}$  and 101.3 kPa pressure.
- f) **SATP** refers to Standard Ambient Temperature and Pressure, and this has been defined as exactly  $25^{\circ}\text{C}$  and 100 kPa pressure. Room conditions are approximately equal to SATP.
- g) A **heating curve** is a graph that shows how temperature changes over time when a substance is heated. In the regions of the graph where the temperature of the substance does not change, the state of the substance is changing (so the potential energy of the particles is increasing).
- h) A **cooling curve** is a graph that shows how temperature changes over time when a substance is cooled. In the regions of the graph where the temperature of the substance does not change, the state of the substance is changing (so the potential energy of the particles is decreasing).
- i) **Boyle's Law** states that the volume of a gas varies inversely with pressure, when the temperature and number of moles of gas are held constant. ( $P_1V_1 = P_2V_2$ )
- j) **Charles' Law** states that the volume of a gas varies directly with temperature (in Kelvins) when the pressure and number of moles of gas are held constant. ( $V_1 / T_1 = V_2 / T_2$ )
- k) **Dalton's Law of Partial Pressure** states that the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures exerted by each of the gases making up the mixture.

### Background for Question #2:

The **Kinetic Molecular Theory**, as it applies to gases, has five basic premises:

- All matter is composed of tiny particles, either atoms or molecules. In the gas state, these particles are far apart relative to their size.
- In the gas state, the particles are in constant random, straight-line motion. The motion follows the normal laws of physics.
- The kinetic energy of these particles depends on the temperature of their surroundings. The higher the temperature the greater the average kinetic energy of the particles.
- The forces of attraction between molecules depend on the inter-molecular attraction between the particles. There is essentially no attraction between the particles of a gas.
- The collisions of gas particles are perfectly elastic. There is no energy lost or gained in these collisions. When the gas particles collide with the sides of the container, they produce pressure. The more collisions that the gaseous particles have with the sides of the container, the higher the pressure of the gas.

# Kinetic Theory Review Answers

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## **Table of Contents Kinetic Theory Review Answers**

1. Understanding the eBook Kinetic Theory Review Answers
  - The Rise of Digital Reading Kinetic Theory Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Kinetic Theory Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Kinetic Theory Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Kinetic Theory Review Answers
  - Personalized Recommendations
  - Kinetic Theory Review Answers User Reviews and Ratings

- Kinetic Theory Review Answers and Bestseller Lists
- 5. Accessing Kinetic Theory Review Answers Free and Paid eBooks
  - Kinetic Theory Review Answers Public Domain eBooks
  - Kinetic Theory Review Answers eBook Subscription Services
  - Kinetic Theory Review Answers Budget-Friendly Options
- 6. Navigating Kinetic Theory Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Kinetic Theory Review Answers Compatibility with Devices
  - Kinetic Theory Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Kinetic Theory Review Answers
  - Highlighting and Note-Taking Kinetic Theory Review Answers
  - Interactive Elements Kinetic Theory Review Answers
- 8. Staying Engaged with Kinetic Theory Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Kinetic Theory Review Answers
- 9. Balancing eBooks and Physical Books Kinetic Theory Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Kinetic Theory Review Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Kinetic Theory Review Answers
  - Setting Reading Goals Kinetic Theory Review Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Kinetic Theory Review Answers
  - Fact-Checking eBook Content of Kinetic Theory Review Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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