

Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction

*iodine
pentoxide*



a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO .

Determine the mass of iodine I_2 , which could be produced?

b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.

i) what mass of iodine was produced?

ii) what percentage yield of iodine was produced.

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

a) Which chemical is the limiting reactant?

b) How many grams of ZnS will be formed?

c) How many grams of the excess reactant will remain after the reaction is over?

3. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen?

What mass is in excess? What mass of MgO is formed?

4. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S ?

5. When MoO_3 and Zn are heated together they react



What mass of ZnO is formed when 20.0 grams of MoO_3 is reacted with 10.0 grams of Zn ?

6. Silver nitrate, AgNO_3 , reacts with ferric chloride, FeCl_3 , to give silver chloride, AgCl , and ferric nitrate, $\text{Fe}(\text{NO}_3)_3$. In a particular experiment, it was planned to mix a solution containing 25.0 g of AgNO_3 with another solution containing 45.0 grams of FeCl_3 .

a) Write the chemical equation for the reaction.

b) Which reactant is the limiting reactant?

c) What is the maximum number of moles of AgCl that could be obtained from this mixture?

Limiting Reagent And Percent Yield 9 3 Review

R Sanford



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Table of Contents Limiting Reagebt And Percent Yield 9 3 Review

1. Understanding the eBook Limiting Reagebt And Percent Yield 9 3 Review
 - The Rise of Digital Reading Limiting Reagebt And Percent Yield 9 3 Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagebt And Percent Yield 9 3 Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagebt And Percent Yield 9 3 Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagebt And Percent Yield 9 3 Review
 - Personalized Recommendations
 - Limiting Reagebt And Percent Yield 9 3 Review User Reviews and Ratings
 - Limiting Reagebt And Percent Yield 9 3 Review and Bestseller Lists
5. Accessing Limiting Reagebt And Percent Yield 9 3 Review Free and Paid eBooks

- Limiting Reagebt And Percent Yield 9 3 Review Public Domain eBooks
- Limiting Reagebt And Percent Yield 9 3 Review eBook Subscription Services
- Limiting Reagebt And Percent Yield 9 3 Review Budget-Friendly Options
- 6. Navigating Limiting Reagebt And Percent Yield 9 3 Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagebt And Percent Yield 9 3 Review Compatibility with Devices
 - Limiting Reagebt And Percent Yield 9 3 Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagebt And Percent Yield 9 3 Review
 - Highlighting and Note-Taking Limiting Reagebt And Percent Yield 9 3 Review
 - Interactive Elements Limiting Reagebt And Percent Yield 9 3 Review
- 8. Staying Engaged with Limiting Reagebt And Percent Yield 9 3 Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagebt And Percent Yield 9 3 Review
- 9. Balancing eBooks and Physical Books Limiting Reagebt And Percent Yield 9 3 Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagebt And Percent Yield 9 3 Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagebt And Percent Yield 9 3 Review
 - Setting Reading Goals Limiting Reagebt And Percent Yield 9 3 Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagebt And Percent Yield 9 3 Review
 - Fact-Checking eBook Content of Limiting Reagebt And Percent Yield 9 3 Review
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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