

**Section C**

Question Number	Acceptable Answers	Mark
<b>24(a)(i)</b>	$(K_c = ) \frac{[\text{CH}_3\text{COOCH}_2\text{CH}_3][\text{H}_2\text{O}]}{[\text{CH}_3\text{COOH}][\text{CH}_3\text{CH}_2\text{OH}]}$ <p><b>ALLOW</b> C<sub>2</sub>H<sub>5</sub>OH for ethanol  <b>ALLOW</b> CH<sub>3</sub>CO<sub>2</sub>H for ethanoic acid  <b>ALLOW</b> CH<sub>3</sub>CO<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub> / CH<sub>3</sub>CO<sub>2</sub>C<sub>2</sub>H<sub>5</sub> / CH<sub>3</sub>COOC<sub>2</sub>H<sub>5</sub> for ethyl ethanoate</p> <p><b>I IGNORE</b> state symbols, even if incorrect</p>	<b>(1)</b>
Question Number	Acceptable Answers	Mark
<b>24(a)(ii)</b>	<p><b>Stand alone marks</b></p> <p>the enthalpy change is (very) small/close to zero  <b>OR</b>  reaction is slightly exothermic <b>(1)</b></p> <p>therefore, (the magnitude of) <math>\Delta S_{\text{surroundings}}</math> (= <math>-\Delta H/T</math>)  <b>changes very little</b> <b>(1)</b></p> <p><b>I IGNORE</b> <math>\Delta S_{\text{surroundings}}</math> is positive/ small/ less/ decreases</p> <p><math>\Delta S_{\text{total}}</math> / <math>K_c</math> changes very little (provided there is no change of state) <b>(1)</b></p> <p>Ignore references to <math>\Delta S_{\text{system}}</math></p>	<b>(3)</b>

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**Clifford Lane**



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