

## Chapter 8 - Chemical Equations and Reactions

### 8-1 Describing Chemical Reactions

#### I. Introduction

##### A. Reactants

1. Original substances entering into a chemical rxn

##### B. Products

1. The resulting substances from a chemical rxn



##### C. Chemical Equation

1. Represents with symbols and formulas, the identities and relative amounts of the reactants and products in a chemical rxn

#### II. Indications of a Chemical Reaction

##### A. Evolution of Heat and Light

1. Evidence of energy being released (exothermic rxn)

##### B. Production of a Gas

1.  $\text{CO}_2$ ,  $\text{H}_2$ ,  $\text{H}_2\text{S}$  are some gases produced by chemical rxns



##### C. Formation of a Precipitate

1. Precipitate is a solid that is produced as a result of a chemical rxn in solution



#### III. Characteristics of Chemical Equations

##### A. The equation must represent known facts

1. This can be done with a word equation

"hydrogen reacts with oxygen to form water"



##### B. The equation must contain the correct formulas for reactants and products

1. This is done with a formula equation



##### C. The law of conservation of atoms must be satisfied

1. Balancing is done with coefficients - small whole numbers that appear in front of a formula



##### D. Additional symbols used in Chemical equations

# Modern Chemistry Chapter 8 Review Answers Section

**Tom Palko**



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**Introduction to Modern Inorganic Chemistry, 6th edition** R.A. Mackay,W. Henderson,2017-12-21 This popular and comprehensive textbook provides all the basic information on inorganic chemistry that undergraduates need to know For this sixth edition the contents have undergone a complete revision to reflect progress in areas of research new and modified techniques and their applications and use of software packages Introduction to Modern Inorganic Chemistry begins by explaining the electronic structure and properties of atoms then describes the principles of bonding in diatomic and polyatomic covalent molecules the solid state and solution chemistry Further on in the book the general properties of the periodic table are studied along with specific elements and groups such as hydrogen the s elements the lanthanides the actinides the transition metals and the p block Simple and advanced examples are mixed throughout to increase the depth of students understanding This edition has a completely new layout including revised artwork case study boxes technical notes and examples All of the problems have been revised and extended and include notes to assist with approaches and solutions It is an excellent tool to help students see how inorganic chemistry applies to medicine the environment and biological topics

*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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