

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that combines internal energy (U) and the product of pressure (P) and volume (V):

$$H = U + PV$$

 - **Enthalpy change** (ΔH) for a process at constant pressure is equal to the heat added to the system:

$$\Delta H = Q_p$$

 - **Enthalpy of formation** (ΔH_f°) is the enthalpy change when one mole of a compound is formed from its elements in their standard states.



- **Gibbs Free Energy** (G) is a thermodynamic potential that measures the maximum reversible work that can be performed by a system at constant temperature and pressure.

$$G = H - TS$$

 where T is the temperature.
 - **Gibbs Free Energy change** (ΔG) for a process at constant temperature and pressure is equal to the maximum non-expansion work that can be extracted from the system:

$$\Delta G = -W_{max}$$



- **Standard Gibbs Free Energy change** (ΔG°) is the Gibbs free energy change for a process under standard conditions (1 bar, 298 K).

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

 where ΔH° is the standard enthalpy change and ΔS° is the standard entropy change.



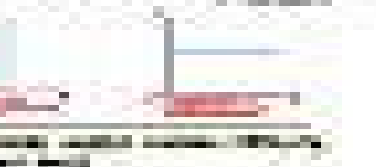
- **Equilibrium** is a state where the macroscopic properties of a system do not change over time.
 - **Chemical Equilibrium** is a state where the concentrations of reactants and products remain constant.

$$K = \frac{[products]}{[reactants]}$$

 where K is the equilibrium constant.



- **Le Chatelier's Principle** states that if a system at equilibrium is subjected to a change in concentration, temperature, or pressure, the system will adjust itself to counteract the change and restore equilibrium.



- **Spontaneous Processes** are those that occur without the need for external energy input.
 - **Non-spontaneous Processes** are those that require external energy input to occur.



Microeconomics Study Guide Final

David M. Wishart



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