#### 2.1 INTRODUCTION

The two main sources of water are

- a) Surface water and
- b) Underground water.

The surface water includes rain water, lake water, river water and sea water. Underground water is further divided into well water and spring water.

#### 2.2 Water quality parameters

#### 2.2.1 Turbidity

Turbidity is the measure of lack of water clarity. Turbidity is the result of suspended solids in water. It can be measured accurately with a turbidimeter. Turbidity is measured in NTUs, i.e. nephelometric turbidity unit. Turbidity in drinking water should be less than 1 NTU. Water treatment is required for drinking water when there is excessive turbidity.

#### 2.2.2 Hardness

Water is said to be hard due to the presence of various dissolved salts of calcium, magnesium and iron. Hardness is expressed as ppm of CaCO<sub>3</sub>. Hardness can be classified into two types, temporary hardness which can be removed by mere boiling and permanent hardness which can be removed by certain water treatment process like lime-soda process, zeolite process etc.

#### 2.2.3 PH

It is in general a measure of hydrogen ion concentration of a water sample. The symbol P<sup>H</sup> stands for the potential of hydrogen. Water contains both H\* ions and OH\* ions. Pure distilled water contains equal number of both the ions and therefore considered to be neutral (P<sup>H</sup> = 7). If a water contains more H\* ions than OH\* ions then the water sample is said to be acidic (P<sup>H</sup> less than 7). If the water contains OH\* ions in excess than H\* ions then the water sample is said to be basic (P<sup>H</sup> is more than 7). Rain water has a P<sup>H</sup> of almost 7 surface water have alkaline P<sup>H</sup>.

#### 2.2.4 Alkalinity

It is the measure of the quantity of alkaline substances in water. Mostly alkalinity in water is due to the presence of hydroxide, carbonate and bicarbonate ions present in the water sample. The possible combinations of ions causing alkalinity are

- (i) OH alone
- (iii) CO<sub>2</sub>2- alone
- (iii) HCO3 alone
- (iv) OH and CO<sub>3</sub>2 together
- (v) CO<sub>3</sub><sup>2</sup> and HCO<sub>3</sub> together.

# **Modern Chemistry Water Review Section 5**

Kenneth Malcolm Mackay,Rosemary
Ann Mackay

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