

# Chemistry: Molality

Directions: Using these formulas, solve the following problems.

$$\text{molality} = \text{mol solute} / \text{kg solvent}$$

$$1 \text{ amu} = 1 \text{ gram} / \text{mol}$$

$$\Delta T_{\text{freezing}} = (- \text{number of ions}) (\text{molality}) (k_f)$$

$$\Delta T_{\text{boiling}} = (\text{number of ions}) (\text{molality}) (k_b)$$

$k_b$  and  $k_f$  are constants for each individual solvent

Key



Problem 1: What is the molality of a solution containing 1.23 g of  $\text{Hg}(\text{CN})_2$  dissolved in 1.37 g of methanol? Answer: **3.55 m**

Problem 2: What is the molality of a solution containing 16.8 g of NaCl dissolved in 356 g of water? Answer: **0.806 m**

Problem 3: How many grams of NaOH would be required to produce a 3.57 m solution with 473 g of water? Answer: **67.6 g**

Problem 4: How many grams of KCl would you need to dissolve in 114 g of vinegar in order to prepare a 1.96 m solution? Answer: **16.6 g**

Problem 5: Calculate the molality of 5.97 g NaCl dissolved in 471 g of water.  
Answer: **0.217 m**

Problem 6: Calculate the molality of 8.91 g  $\text{Na}_2\text{CO}_3$  dissolved in 595 g of water.  
Answer: **0.341 m**

Problem 7: Calculate the molality of 8.44 g KBr dissolved in 112 g of water.  
Answer: **0.633 m**

Problem 8: What are the new boiling and freezing points for a solution with 581 g KBr dissolved in 845 grams of water? ( $k_b$  water =  $0.512^\circ \text{C/m}$  and  $k_f$  water =  $1.86^\circ \text{C/m}$ )  
Answer: new freezing point =  **$-1.34^\circ \text{C}$**       new boiling point =  **$106^\circ \text{C}$**

Problem 9: What are the new boiling and freezing points for a solution with 291 g NaOH dissolved in 1029 grams of water? ( $k_b$  water =  $0.512^\circ \text{C/m}$  and  $k_f$  water =  $1.86^\circ \text{C/m}$ )  
Answer: new freezing point =  **$-36.3^\circ \text{C}$**       new boiling point =  **$107^\circ \text{C}$**

# Molarity And Molality Worksheet

**Mary Louise Turgeon**



## **Molarity And Molality Worksheet:**

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*Holt Chemistry* Holt Rinehart & Winston, 2003-01-24

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**AIChE Symposium Series** American Institute of Chemical Engineers, 1972

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