

Dilution Calculations

What is the molarity (M) of a solution prepared by diluting 0.180L of 0.600 M HNO_3 to 0.540 L?

Prepare a table:

$$M_1 = 0.600 \text{ M} \quad V_1 = 0.180 \text{ L}$$

$$M_2 = ? \quad V_2 = 0.540 \text{ L}$$

Solve dilution equation for unknown and enter values:

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2} = \frac{(0.600 \text{ M})(0.180 \text{ L})}{0.540 \text{ L}} = 0.200 \text{ M}$$

Molarity By Dilution Chemistry If8766

R Bogdan



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Molarity By Dilution Chemistry If8766 Book Review: Unveiling the Power of Words

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