

## Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of  $\text{N}_2\text{O}_4$  was produced, how many moles of  $\text{O}_2$  were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1\text{ mol N}_2\text{O}_4}{1\text{ mol N}_2\text{O}_4} \times \frac{2\text{ mol O}_2}{1\text{ mol N}_2\text{O}_4} = 0.326\text{ mol O}_2$$

- b. If  $4.0 \times 10^{-3}$  moles of oxygen reacted, how many grams of  $\text{N}_2$  were needed?

$$\frac{4.0 \times 10^{-3}\text{ mol O}_2}{2\text{ mol O}_2} \times \frac{1\text{ mol N}_2}{1\text{ mol O}_2} \times \frac{28.0\text{g N}_2}{1\text{ mol N}_2} = 5.6 \times 10^{-2}\text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04\text{ mol K}_3\text{PO}_4}{1\text{ mol K}_3\text{PO}_4} \times \frac{3\text{ mol KNO}_3}{1\text{ mol K}_3\text{PO}_4} \times \frac{101.1\text{ g KNO}_3}{1\text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1\text{ mol AlPO}_4}{1\text{ mol AlPO}_4} \times \frac{1\text{ mol Al}(\text{NO}_3)_3}{1\text{ mol AlPO}_4} = 4.75 \times 10^{-2}\text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of  $\text{CaC}_2$ , how much  $\text{C}_2\text{H}_2$  do you get?

$$\frac{5.50\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{1\text{ mol C}_2\text{H}_2}{1\text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of  $\text{CaC}_2$  have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{2\text{ mol H}_2\text{O}}{1\text{ mol CaC}_2} = 2.03\text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ). Write and balance the chemical equation. How many moles of  $\text{CO}_2$  are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0\text{g C}_6\text{H}_{12}\text{O}_6} \times \frac{1\text{ mol C}_6\text{H}_{12}\text{O}_6}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6\text{ mol CO}_2}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000\text{ mol CO}_2$$

# Mole Lab Answer Key

**Karen C. Timberlake**



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Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books The teacher is one who guides students through the subject matter helps each student stay on schedule and be organized and is their source of accountability along the way With that in mind this guide provides additional help through the laboratory exercises as well as lessons quizzes and examinations that are provided along with the answers The lessons in this study emphasize working through procedures and problem solving by learning patterns The vocabulary is kept at the essential level Practice exercises are given with their answers so that the patterns can be used in problem solving These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college Guided labs are provided to enhance instruction of weekly lessons There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions It is important to see the hand of God and His principles and wisdom as it plays out in chemistry This course integrates what God has told us in the context of this study Features Each suggested weekly schedule has five easy to manage lessons that combine reading and worksheets Worksheets quizzes and tests are perforated and three hole punched materials are easy to tear out hand out grade and store Adjust the schedule and materials needed to best work within your educational program Space is given for assignments dates There is flexibility in scheduling Adapt the days to your school schedule Workflow Students will read the pages in their book and then complete each section of the teacher guide They should be encouraged to complete as many of the activities and projects as possible as well Tests are given at regular intervals with space to record each grade About the Author DR DENNIS ENGLIN earned his bachelor's from Westmont College his master of science from California State University and his EdD from the University of Southern California He enjoys teaching animal biology vertebrate biology wildlife biology organismic biology and astronomy at The Master's University His professional memberships include the Creation Research Society the American Fisheries Association Southern California Academy of Sciences Yellowstone Association and Au Sable Institute of Environmental Studies Chemistry Greg Curran, 2011 Covers all the topics in a typical one year high school chemistry curriculum

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