

Mole Conversions Worksheet

Solve the following; express your answer using the proper amount of significant figures.

1. How many moles of magnesium are in 3.01×10^{22} atoms of magnesium?
0.0500 moles of magnesium
2. How many molecules are there in 4.00 moles of glucose, $C_6H_{12}O_6$?
 2.41×10^{24} molecules of glucose
3. How many moles are 1.20×10^{25} atoms of phosphorous?
19.9 moles of phosphorous
4. How many moles are in 28 grams of CO_2 ?
0.64 moles of carbon dioxide
5. What is the mass of 5 moles of Fe_2O_3 ?
 8×10^2 grams of iron (III) oxide
6. Find the number of moles of argon in 452 g of argon.
11.3 moles of Argon
7. Find the mass in grams of 2.00×10^{23} molecules of F_2 .
12.6 grams of difluorine
8. Find the mass, in grams, of 1.00×10^{23} molecules of N_2 .
4.65 grams of dinitrogen
9. How many particles are there in 1.43 g of a molecular compound with a molar mass of 233 g? **3.70×10^{21} particles**
10. Aspartame is an artificial sweetener that is 160 times sweeter than sucrose (table sugar) when dissolved in water. The molecular formula of aspartame is $C_{14}H_{18}N_2O_5$.
 - a. Calculate the molar mass of aspartame.
294.34 grams/mole
 - b. How many moles are in 10 g of aspartame?
0.03 moles of aspartame
 - c. How many molecules are in 5 mg of aspartame?
 1×10^{19} molecules of aspartame
 - d. What is the mass, in grams, of 1.56 moles of aspartame?
459 grams of aspartame
 - e. How many atoms of nitrogen are in 1.2 grams of aspartame?
 4.9×10^{21} atoms of nitrogen

Mole A Measurement Of Matter Answer Key

Scott C. Dulebohn



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