

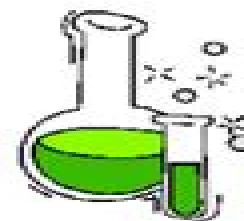
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The Mole Lab

Chemistry I Acc

(Weighing as a Means of Counting)



Introduction

One of the seven SI base units is the mole. The mole, also known as Avogadro's number, is equal to 6.02×10^{23} . The mole is a quantity like a dozen (12) or a gross (144). If you wanted to know how many eggs were in 3 dozen eggs you would multiply 3 dozen eggs \times 12 eggs/dozen. If you wanted to know how many firecrackers were in 3 gross firecrackers you would multiply 3 gross firecrackers \times 144 firecrackers/gross. Likewise if you wanted to know how many atoms were in 3 moles of atoms you would multiply 3 moles of atoms \times 6.02×10^{23} atoms/mole.

The Atomic Mass Unit is defined as 1/12 of a Carbon-12 atom. It is the standard in the AMU scale. The mass of all atoms in AMUs is defined relative to this standard. Likewise, the mass of a mole is defined as the number in grams that is equal to the average AMU of the element. For example the average atomic mass of Carbon is 12.011 AMU. The mass of one mole of Carbon is equal to 12.011 grams. The mass of one mole of Lead is equal to 207.2 grams. The mass of one mole of a molecule is equal to the mass of one mole of each of the atoms in the compound. For example the mass of one mole of Silicon dioxide is equal to 28.0855 grams for the Silicon plus 15.9994 grams for each of the oxygens (To make our lives easy we will round masses of all elements to the nearest 0.1 grams). Therefore, mass of one mole of SiO_2 is equal to 60.1 grams ($28.1 + 16.0 + 16.0$ grams). The mass of one mole of substance (element or compound) is called the substance's molar mass in units of grams/mole.

In part one of this lab you will convert between mass, moles and atoms of a given element and identify a substance based on calculating its molar mass. In part two you will use measurements of mass in order to convert between mass, moles, and molecules or atoms.

Materials

Part One:

Nine vials of various elements
Triple Beam Balance

Part Two:

1 canister Sodium Chloride
1 canister Calcium Carbonate
1 2 oz. Cup for Water
1 weighing boat for Sodium Chloride
1 weighing boat for Calcium Carbonate
1 stirrer
2 spoons

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N Colangelo



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