

Moorpark College
Chemistry 1A Mock Exam #1 **ANSWER KEY**
Professor Torres

1. A chemistry student has the idea of removing a decorative stone sphere with a radius of 0.289 m from the front of a campus building. If the density of the stone is 3.52 g/cm^3 , what is the mass of the sphere?

$$V_{\text{sphere}} = \frac{4}{3} \pi r^3$$

$$\text{Volume} = \frac{4}{3} \pi r^3 = \frac{4}{3} \times \pi \times (28.9 \text{ cm})^3 = 1.01 \times 10^5 \text{ cm}^3$$

$$\text{Density} = \text{mass/volume} \Rightarrow \text{Mass} = \text{Density} \times \text{Volume}$$

$$\text{Mass} = (3.52 \text{ g/cm}^3)(1.01 \times 10^5 \text{ cm}^3) = \boxed{3.56 \times 10^5 \text{ g}}$$

2. Determine the mass (in grams) of 1.00 gal of cyclohexane. The density of cyclohexane at 25°C is 0.7781 g/mL . [NOTE: $1 \text{ L} = 1.057 \text{ qt}$ and $1 \text{ gal} = 4 \text{ qt}$]

$$\begin{aligned} \text{Mass (g)} &= 1.00 \text{ gal} \times \frac{4 \text{ qt}}{1 \text{ gal}} \times \frac{1 \text{ L}}{1.057 \text{ qt}} \times \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{0.7781 \text{ g}}{\text{mL}} \\ &= \boxed{2.94 \times 10^3 \text{ g cyclohexane}} \end{aligned}$$

3. CFC's, or chlorofluorocarbons, are damaging pollutants to the earth's atmosphere that are known as greenhouse gases. Above 30,000 feet, the atmosphere has a density of $1.2 \times 10^{-6} \text{ g/cm}^3$. Determine the mass of CFCl_3 (a CFC) contained in $7.2 \times 10^6 \text{ cm}^3$ air if the CFCl_3 concentration is 50. ppm (i.e. one gram of CFCl_3 in one million grams of air).

$$\begin{aligned} \text{grams CFCl}_3 &= \frac{50 \text{ g CFCl}_3}{10^6 \text{ g air}} \times \frac{1.2 \times 10^{-6} \text{ g air}}{\text{cm}^3} \times 7.2 \times 10^6 \text{ cm}^3 \text{ air} \\ &= \boxed{4.32 \times 10^{-1} \text{ g CFCl}_3} \end{aligned}$$

4. The element potassium has two naturally occurring isotopes: ^{39}K and ^{41}K with a mass of 40.962 amu. Potassium consists of 6.88% ^{41}K and has an average atomic mass of 39.101 amu. Calculate the mass of ^{39}K .

$$\begin{aligned} x &= \text{mass of } ^{39}\text{K} \\ 39.101 &= 40.962(0.0688) + 0.9312x \\ x &= \boxed{38.964 \text{ amu}} \end{aligned}$$

5. Nitrogen has two naturally occurring isotopes, ^{14}N and ^{15}N , whose atomic masses are 14.0031 and 15.0001, respectively. The atomic mass of nitrogen is 14.0067. Calculate the percentage abundance of each isotope of nitrogen.

$$\begin{aligned} x &= \% \text{ abundance of } ^{14}\text{N}; \quad 1 - x = \% \text{ abundance } ^{15}\text{N} \\ 14.0067 &= 14.0031x + 15.0001(1 - x) \\ x &= 0.9964 \text{ or } \boxed{99.64\% \text{ } ^{14}\text{N}; \text{ therefore, } 0.36\% \text{ } ^{15}\text{N}} \end{aligned}$$

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