

Name: \_\_\_\_\_

## Balancing Chemical Equations

Directions: Balance each of the following chemical equations.



- 1)    \_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_  $\text{KOH}$   $\rightarrow$  \_\_\_  $\text{K}_3\text{PO}_4$  + \_\_\_  $\text{H}_2\text{O}$
- 2)    \_\_\_  $\text{K}$  + \_\_\_  $\text{B}_2\text{O}_3$   $\rightarrow$  \_\_\_  $\text{K}_2\text{O}$  + \_\_\_  $\text{B}$
- 3)    \_\_\_  $\text{HCl}$  + \_\_\_  $\text{NaOH}$   $\rightarrow$  \_\_\_  $\text{NaCl}$  + \_\_\_  $\text{H}_2\text{O}$
- 4)    \_\_\_  $\text{Na}$  + \_\_\_  $\text{NaNO}_3$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$  + \_\_\_  $\text{N}_2$
- 5)    \_\_\_  $\text{C}$  + \_\_\_  $\text{S}_8$   $\rightarrow$  \_\_\_  $\text{CS}_2$
- 6)    \_\_\_  $\text{Na}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$
- 7)    \_\_\_  $\text{N}_2$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{N}_2\text{O}_5$
- 8)    \_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_  $\text{Mg}(\text{OH})_2$   $\rightarrow$  \_\_\_  $\text{Mg}_3(\text{PO}_4)_2$  + \_\_\_  $\text{H}_2\text{O}$
- 9)    \_\_\_  $\text{NaOH}$  + \_\_\_  $\text{H}_2\text{CO}_3$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{CO}_3$  + \_\_\_  $\text{H}_2\text{O}$
- 10)   \_\_\_  $\text{KOH}$  + \_\_\_  $\text{HBr}$   $\rightarrow$  \_\_\_  $\text{KBr}$  + \_\_\_  $\text{H}_2\text{O}$
- 11)   \_\_\_  $\text{Na}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$
- 12)   \_\_\_  $\text{Al}(\text{OH})_3$  + \_\_\_  $\text{H}_2\text{CO}_3$   $\rightarrow$  \_\_\_  $\text{Al}_2(\text{CO}_3)_3$  + \_\_\_  $\text{H}_2\text{O}$
- 13)   \_\_\_  $\text{Al}$  + \_\_\_  $\text{S}_8$   $\rightarrow$  \_\_\_  $\text{Al}_2\text{S}_3$
- 14)   \_\_\_  $\text{Cs}$  + \_\_\_  $\text{N}_2$   $\rightarrow$  \_\_\_  $\text{Cs}_3\text{N}$
- 15)   \_\_\_  $\text{Mg}$  + \_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_  $\text{MgCl}_2$
- 16)   \_\_\_  $\text{Rb}$  + \_\_\_  $\text{RbNO}_3$   $\rightarrow$  \_\_\_  $\text{Rb}_2\text{O}$  + \_\_\_  $\text{N}_2$
- 17)   \_\_\_  $\text{C}_6\text{H}_6$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$  + \_\_\_  $\text{H}_2\text{O}$
- 18)   \_\_\_  $\text{N}_2$  + \_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_  $\text{NH}_3$

# Math Skills Balancing Chemical Equations Answers

**CL Gary**



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