General Chemistry 1 Quarter 1, Lesson 14 LIMITING AND EXCESS REAGENTS

Objectives:

- Define the concept of limiting and excess reactants.
- Determine the limiting and excess reactants using mole ratios from balanced chemical equations.

When a chemical reaction is carried out in a flask, the amount of reactants are not always present in the exact proportion stoichiometrically determined from the balanced equation. To make sure that the more expensive reagent is completely used up and converted to the desired product, chemists usually add the cheaper reagent in excess quantity. The reagent that has completely reacted and used up in a reaction is called the **limiting reagent**. The **excess reagent** is the reactant that is present in quantity higher than what is required to react with the limiting reagent.

Sample Problem 1. Let's use the chemical equation below:
$$3H_2(g) + N_2(g) \rightarrow 2NH_1(g)$$

Ammonia, NH_3 , is synthesized from the reaction of H_2 and N_2 gases. Suppose 6 moles of H_2 was initially mixed with 4 moles of N_2 gas at high pressure. To determine which of the 2 reactants is the limiting reagent, the amount of NH_3 produced must be computed given the number of moles of H_2 and N_2 and the mole-mole factor from the balanced equation.

moles
$$NH_3 = \# \text{ moles of } H_2 \times \frac{2 \text{ moles } NH_3}{3 \text{ moles } H_2}$$

moles $NH_3 = \# \text{ moles of } H_2 \times \frac{2 \text{ moles } NH_3}{3 \text{ moles } H_2}$

moles $NH_3 = 6 \text{ moles } Of H_2 \times \frac{2 \text{ moles } NH_3}{3 \text{ moles } H_2}$

moles $NH_3 = 4 \text{ moles } Of H_2 \times \frac{2 \text{ moles } NH_3}{4 \text{ mole } NH_3}$

moles $NH_3 = 4 \text{ moles } Of H_2 \times \frac{2 \text{ moles } NH_3}{4 \text{ mole } NH_3}$

moles $NH_3 = 6 \text{ moles } NH_3 = 6 \text{ moles } NH_3$

moles $NH_3 = 6 \text{ moles } NH_3 = 6 \text{ moles } NH_3$

From the computed values we determined that if 6 moles of H_2 completely reacts with N_2 , it can produce 4 moles of NH_2 while 4 moles of N_2 can produce 8 moles of NH_2 when fully used up. Since there is less amount of NH_2 produced with 6 moles of H_2 than 4 moles of N_2 , H_2 gas is the limiting reagent while the N_2 gas is the reagent in excess. To determine how much of the 4 moles of N_2 is in excess, we will use the mole-mole factor of N_2 and H_2 .

moles
$$N_2 = 6$$
 moles of $H_2 \times \frac{1 \text{ mole } N_2}{3 \text{ moles } H_2}$ The number of moles of N_2 required to fully react to 6 moles of H_2 is only 2 moles. Thus, the initial 4 moles of N_2 has an excess of 2 moles.

$$excess N_2 = 2 \text{ moles} - 2 \text{ moles} = 2 \text{ moles}$$

Sample Problem 2. The combustion of ethane produces carbon dioxide and water shown in the reaction below:

$$2C_2H_6 + 7O_2 \rightarrow 4CO_2 + 6H_2O$$

 How many moles of CO₂ are produced with 56.2 moles C₂H₆ and 73.4 moles of O₂? Solution:

moles
$$CO_2 = 56.2$$
 moles of $C_2H_4 \times \frac{4 \text{ moles } CO_2}{5 \text{ moles } CO_2}$ moles $CO_2 = 73.4$ moles $CO_2 \times \frac{4 \text{ moles } CO_2}{7 \text{ moles } CO_2}$ moles $CO_2 = 41.9$ moles $CO_3 \times \frac{4 \text{ moles } CO_2}{7 \text{ moles } CO_3}$

- Which reagent is the limiting reagent? Which is the excess reagent?
 The Limiting reagent is O₂ while the excess reagent is C₂H₆.
- How many moles of C₂H₆ are in excess?

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- 1. Understanding the eBook Limited And Excess Reactants Pogil Answer Key
 - The Rise of Digital Reading Limited And Excess Reactants Pogil Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Limited And Excess Reactants Pogil Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limited And Excess Reactants Pogil Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Limited And Excess Reactants Pogil Answer Key
 - Personalized Recommendations
 - Limited And Excess Reactants Pogil Answer Key User Reviews and Ratings
 - Limited And Excess Reactants Pogil Answer Key and Bestseller Lists

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 - Limited And Excess Reactants Pogil Answer Key eBook Subscription Services
 - Limited And Excess Reactants Pogil Answer Key Budget-Friendly Options
- 6. Navigating Limited And Excess Reactants Pogil Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
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 - Limited And Excess Reactants Pogil Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limited And Excess Reactants Pogil Answer Key
 - Highlighting and Note-Taking Limited And Excess Reactants Pogil Answer Key
 - Interactive Elements Limited And Excess Reactants Pogil Answer Key
- 8. Staying Engaged with Limited And Excess Reactants Pogil Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limited And Excess Reactants Pogil Answer Key
- 9. Balancing eBooks and Physical Books Limited And Excess Reactants Pogil Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limited And Excess Reactants Pogil Answer Key
- 10. Overcoming Reading Challenges
 - $\circ\,$ Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limited And Excess Reactants Pogil Answer Key
 - Setting Reading Goals Limited And Excess Reactants Pogil Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limited And Excess Reactants Pogil Answer Key
 - Fact-Checking eBook Content of Limited And Excess Reactants Pogil Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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