Limiting Reactant and Percentage Yield

1 Calcium hydroxide reacts with hydrochloric acid according to the following reaction:

Ca(OH), + 2 HCl -> CaCl, + 2 H,O

(a) Suppose 6.3 mol of HCl reacts with 2.8 mol Ca(OH)₂. Which substance is the limiting reactant?

- (b) How many moles of the excess reactant remain?
- 2 Aluminum oxidizes according to the following unbalanced reaction: (First, balance it)

__ Al + __ O, → __ Al,O,

a 0.048 mol of powdered Al is placed in a container containing 0.030 mol of O₂. What is the limiting reactant?

- (b) How many moles of excess reactant remain?
- 3 In the reaction between CO and Fe₃O_a, the theoretical yield in an experiment is calculated to be 47.2 g Fe. When a careless chemistry student experiments, the actual yield is 42.9 g Fe. Calculate the percentage yield.

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Davis, Peck, STANLE

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