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Emilie Sanchez

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 time the total pressure $p_a = x_a p_{\text{total}}$ mole fraction $x_a = \frac{\text{moles of substance a}}{\text{total moles}}$
 a is often denoted by x_a and

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 present is equal to one $x_a + x_b = 1$ when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is
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 ways to represent the concentration of the solution it is defined as the number of molecules of component a divided by the
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 mass $C_6H_6 = 78 \text{ g/mol}$ $12 \times 6 + 1 \times 6 = 78 \text{ g/mol}$ hence C_6H_6 's number of moles $\frac{30}{78} = 0.3846 \text{ mol}$ molar mass of carbon tet $CCl_4 = 154 \text{ g/mol}$
 $12 \times 4 + 35.5 \times 4 = 154 \text{ g/mol}$

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 g of water calculate the molarity and molality of the solution given $N = 14$ $H = 1$ $C = 12$ $O = 16$ given mass of solute urea 11.11 g mass
 of solvent water 100 g 0.1 kg to find molarity of solution solution

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 of $CH_3OH = \frac{0.173}{0.173 + 2.23} = 0.073$ to solve more examples on mole fraction formula and practice more
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