Limiting Reagent and Percentage Yield

Dichlorine monoxide is produced by passing chlorine gas over heated mercury (II) oxide according to the following unbalanced reaction:

$$_$$
 HgO \longrightarrow $_$ Cl₂ + $_$ HgCl₂ + $_$ Cl₂O

Balance the reaction. Suppose the quantity of the reactants is sufficient to produce 0.86 g of Cl₂O, but only 0.71 g is obtained. What is the percentage yield?

2 Consider the following reaction:

The theoretical yield of beryllium chloride was 10.7 g. Suppose the reaction actually yields 4.5 g. What is the percentage yield?

3 Nitrogen and hydrogen at high temperatures are converted to ammonia using the following reaction:

$$N_2 + 3 H_3 \longrightarrow 2 NH_3$$

When 400 g of H_2 are added to an excess amount of N_2 , 104 g of NH_3 are formed. Calculate the percentage yield.

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Angela Woodward Spangenberg

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