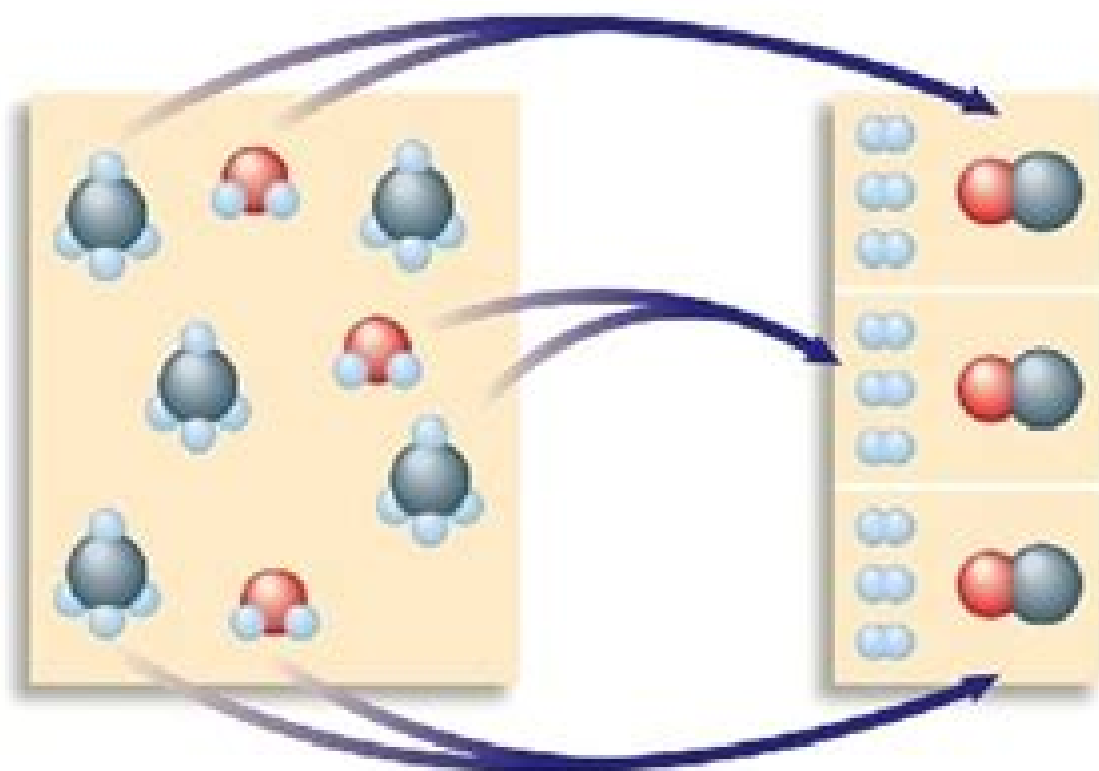


Limiting Reactants and Percent Yield

B. Calculations Involving a Limiting Reactant

A mixture of 5 CH_4 and 3 H_2O molecules undergoes the reaction
 $\text{CH}_4(g) + \text{H}_2\text{O}(g) \rightarrow 3\text{H}_2(g) + \text{CO}(g)$.
Note that the H_2O molecules are used up first, leaving two
 CH_4 molecules unreacted.



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John Moore, Richard H. Langley



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Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

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