

Chemistry 12 – Unit 1 – Reaction Kinetics
Kinetics Review Test

1. Consider the following reaction mechanism:



The chemical species MX is a(n) _____.

- A. catalyst B. inhibitor C. product D. reaction intermediate
2. Consider the following reaction:

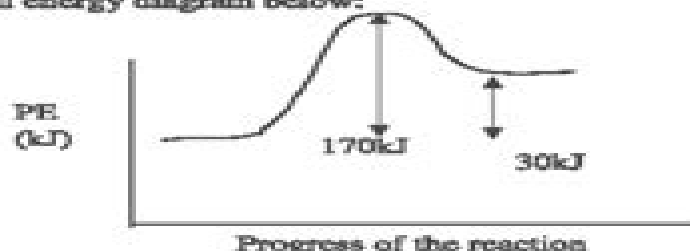
$$2 \text{N}_2\text{O}_5(g) \rightarrow 4 \text{NO}_2(g) + \text{O}_2(g)$$
 At a certain temperature the rate of decomposition of $\text{N}_2\text{O}_{5(g)}$ is $2.5 \times 10^{-6} \text{ mol/s}$. The rate of formation of NO_2 is _____.

- A. $1.0 \times 10^{-5} \text{ mol/s}$ B. $1.3 \times 10^{-6} \text{ mol/s}$ C. $2.5 \times 10^{-6} \text{ mol/s}$ D. $5.0 \times 10^{-6} \text{ mol/s}$
3. Which of the following factors affect the rates of *both* homogeneous and heterogeneous reactions?
- | | |
|--------------------------------|-----------------------------------|
| I nature of the reactants | II presence of a catalyst |
| III temperature of system | IV concentration of reactants |
- A. I and IV only B. II and III only C. II, III, and IV only D. I, II, III, and IV

4. Which of the following equations represents an endothermic reaction?



5. Consider the potential energy diagram below:



The activation energy for the reverse reaction is _____.

- A. 30 kJ B. 140 kJ C. 170 kJ D. 200 kJ

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