

CHAPTER

06

THERMODYNAMICS

- + Universe = system + surroundings
- + $W = F \times dx$ where W = work done, F = force, dx = distance moved
- + $W = P_{ext} \Delta V$ ($\Delta V = V_2 - V_1$)
- + $dW = P dV$
- + $W = \int_{V_1}^{V_2} P_{ext} dV$
- + Work done by a system = -ve
- + Work done on the system = +ve
- + $C = \frac{q}{dT}$; q = Heat absorbed, dT = rise in temperature
- + $Q = \Delta E + W$ I law of thermodynamics
- + $q = dE + W$ I law of thermodynamics
- + Heat absorbed by a system = +ve
- + Heat given out by a system = -ve
- + $H = E + PV$
- + $\Delta H = \Delta E + P \Delta V$
- + $C_v = \left(\frac{\partial E}{\partial T} \right)_v$; C_v = Heat capacity at constant volume
 ∂E = Change in internal energy, ∂T = change in temperature at constant volume in Kelvin
- + $C_p = \left(\frac{\partial H}{\partial T} \right)_p$
- + $C_p - C_v = R$, R = Gas constant, C_p = Heat capacity at constant pressure
- + ∂H = Change in enthalpy, ∂T = Change in temperature at constant pressure
- + $\frac{C_p}{C_v} = \gamma$
- + Molar heat capacity = specific heat x molecular weight
- + Heat liberating reactions are exothermic reactions, $\Delta H = -ve$
- + Heat absorbing reactions are endothermic reactions, $\Delta H = +ve$
- + Heat of combustion = $Z \times \theta \times \frac{M}{m} \text{ (cals) }$,
 M = molecular weight, m = weight of substance
 θ = rise in temperature
 Z = heat capacity
- + Heat liberated = (W + volume of reaction mixture) x rise in temperature
- + $W = \left[\frac{m_2 (t_2 - t_1)}{t_3 - t_1} - m_1 \right]$, m_1 = mass of water at lower temperature
 m_2 = mass of water at higher temperature
 $t_1^\circ C$ = Lower temperature, $t_2^\circ C$ = higher temperature
 $t_3^\circ C$ = resultant temperature, W = water equivalent of calorimeter + stirrer

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RS Peters

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web the court however proceeded to effectively disregard section 27 3 by looking to section 59 1 a ii which allows the minister to make regulations necessary for the effective implementation and enforcement generally of the act s provisions

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