

- **Balmer Series.** When excited electrons in hydrogen atoms fall from higher energy levels to second energy level, the series of lines observed is called Balmer series. They are observed in visible region.

$$\bar{n} = R \left[\frac{1}{2^2} - \frac{1}{n^2} \right], n = 3, 4, 5, \dots$$

- **Paschen Series.** When excited electrons in hydrogen atoms fall from higher energy levels to third energy level, we observe Paschen series. They are observed in infra-red region.

$$\text{Paschen } \bar{n} = R \left[\frac{1}{3^2} - \frac{1}{n^2} \right], n = 4, 5, 6, 7, \dots$$

- **Brackett Series.** When excited electrons in hydrogen atoms fall from higher energy levels of fifth energy level, we observe Brackett series. They are also observed in infra-red region of spectrum.

Brackett

$$\bar{n} = R \left[\frac{1}{4^2} - \frac{1}{n^2} \right], n = 5, 6, 7, 8, \dots$$

- **Pfund Series.** When excited electrons in hydrogen atoms fall from higher energy levels of fifth energy level, Pfund series is observed. They are also found in infra-red region of spectrum.

$$\text{Pfund } \bar{n} = R \left[\frac{1}{5^2} - \frac{1}{n^2} \right], n = 6, 7, 8, 9, \dots$$

- **Rydberg Formula.**

$$\bar{n} = \frac{1}{\lambda} = R_H \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$

Where \bar{n} is wave number, λ is wavelength, R_H is Rydberg constant, n_1 is lower energy level, n_2 is higher energy level, $R_H = 109679 \text{ cm}^{-1}$.

- **Bohr's Theory:**

- (i) An electron did not radiate energy if it stayed in one orbit and therefore did not slow down

- (ii) When an electron moved from one orbit to another, it either radiates or absorbs energy. If it moved towards nucleus, energy is radiated.

$$\Delta E = E_2 - E_1 = h\nu = \frac{hc}{\lambda}$$

$$E_n = \frac{-1312}{n^2} \text{ kJ mol}^{-1} \text{ for H-atom}$$

$$E_n = \frac{-2\pi^2 m Z^2 e^4}{n^2 h^2} \text{ for other atom}$$

$$E_n = \frac{-1312 \cdot 10^3 \text{ J mol}^{-1}}{n^2} \text{ for H-atom}$$

$$E_n = \frac{-2.173 \cdot 10^{-18} \text{ J/atom}}{n^2}$$

For Hydrogen like species

$$E_n = \frac{-1312 \cdot Z^2 \text{ kJ mol}^{-1}}{n^2},$$

where 'Z' is atomic number.

$$\Delta E = E_2 - E_1 = \frac{hc}{\lambda} = 1312 \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right] \text{ kJ mol}^{-1}$$

$$r_n = \frac{n^2 h^2}{4\pi^2 m e^2 Z} \text{ or } r_n \propto n^2 \text{ or } r_n \propto \frac{1}{Z}$$

$r_n = 0.0529 n^2 \text{ nm}$ for H-atom, where 'n' is energy level, r_n is radius of nth orbit.

$$r_n = \frac{0.0529 n^2}{Z}, \text{ where 'Z' is atomic number.}$$

(For H-like atom)

Radius of various nuclei = $1.4 \times 10^{-13} \text{ cm} \times A^{1/3}$

$$V_n = \frac{2\pi e^2 Z}{nh} = \frac{2.188 \cdot 10^6 \cdot Z}{n} \text{ ms}^{-1}$$

- (iii) The angular momentum of orbits in which electron revolves must be integral multiple

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M Planty



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Table of Contents Modern Chemistry Ch 7 Section 4review Answers

1. Understanding the eBook Modern Chemistry Ch 7 Section 4review Answers
 - The Rise of Digital Reading Modern Chemistry Ch 7 Section 4review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Ch 7 Section 4review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Ch 7 Section 4review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Ch 7 Section 4review Answers
 - Personalized Recommendations
 - Modern Chemistry Ch 7 Section 4review Answers User Reviews and Ratings
 - Modern Chemistry Ch 7 Section 4review Answers and Bestseller Lists
5. Accessing Modern Chemistry Ch 7 Section 4review Answers Free and Paid eBooks
 - Modern Chemistry Ch 7 Section 4review Answers Public Domain eBooks
 - Modern Chemistry Ch 7 Section 4review Answers eBook Subscription Services
 - Modern Chemistry Ch 7 Section 4review Answers Budget-Friendly Options
6. Navigating Modern Chemistry Ch 7 Section 4review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Ch 7 Section 4review Answers Compatibility with Devices
 - Modern Chemistry Ch 7 Section 4review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Ch 7 Section 4review Answers
 - Highlighting and Note-Taking Modern Chemistry Ch 7 Section 4review Answers
 - Interactive Elements Modern Chemistry Ch 7 Section 4review Answers
8. Staying Engaged with Modern Chemistry Ch 7 Section 4review Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Modern Chemistry Ch 7 Section 4review Answers
- 9. Balancing eBooks and Physical Books Modern Chemistry Ch 7 Section 4review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Ch 7 Section 4review Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Ch 7 Section 4review Answers
 - Setting Reading Goals Modern Chemistry Ch 7 Section 4review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Ch 7 Section 4review Answers
 - Fact-Checking eBook Content of Modern Chemistry Ch 7 Section 4review Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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