

## Chapter Test A

### Chapter: Atoms: The Building Blocks of Matter

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. Neutral atoms contain equal numbers of  
a. electrons and neutrons.  
b. protons and neutrons.  
c. protons and electrons.  
d. protons, electrons, and neutrons.
- \_\_\_\_\_ 2. The atomic symbol for beryllium,  ${}^9_4\text{Be}$ , indicates that the  
a. atomic number is 4.  
b. atomic number is 9.  
c. mass number is 4.  
d. atomic number is equal to  $9 - 4$ .
- \_\_\_\_\_ 3. The atomic mass of an atom of carbon is 12 amu, and the atomic mass of an atom of oxygen is 16 amu. To produce CO, 16 g of oxygen combine with 12 g of carbon. To produce  $\text{CO}_2$ , 32 g of oxygen combine with 12 g of carbon. The ratio of the mass of oxygen in  $\text{CO}_2$  to the mass of oxygen in CO is  
a. 1:1.  
b. 2:1.  
c. 1:2.  
d. 8:3.
- \_\_\_\_\_ 4. Rutherford's gold-foil experiment led him to conclude that  
a. Thomson's plum pudding model of the atom was accurate.  
b. alpha particles were a poor choice for a bombardment material.  
c. a dense region of positive charge existed somewhere in the atom.  
d. light was emitted by electrons returning to ground state.
- \_\_\_\_\_ 5. Which of the following is *not* one of the five principles of Dalton's theory?  
a. Atoms of different elements combine in simple, whole-number ratios to form compounds.  
b. All matter is made of indivisible, indestructible atoms.  
c. All atoms have similar physical and chemical properties.  
d. Chemical reactions consist of the combination, separation, or rearrangement of atoms.

# Modern Chemistry Chapter Review Atoms Building Blocks Matter Answers

**Rachel Sandford**



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